

Drinking Water Plant Operator Certification Training



Module 25: Hypochlorite

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Appendix

SDS for sodium hypochlorite

Additional Hypochlorite Practice Math Problems

Unit 1 – Background and Properties

Learning Objectives

- Outline the history of hypochlorite use.
- List the uses of hypochlorite.
- Explain how hypochlorite is produced.
- List and explain six properties of hypochlorite.

History of Use

- During the early history of water chlorination, the only sources of chlorine were dry chlorine-containing compounds and sodium hypochlorite bleach solutions. Initially, the following problems occurred:
 - The poor stability and variable effective chlorine content resulted in operating difficulties.
 - Feeder equipment was crude, therefore, yielding erratic results.
- In 1909, liquid chlorine became commercially available and the use of hypochlorite for water chlorination gradually decreased in popularity.
- In 1928, the commercial availability of calcium hypochlorite resulted in renewed use.

Uses



Hypochlorite refers to the various salts of hypochlorous acid commonly used in water treatment for disinfection, oxidation, and taste and odor control. This term is commonly used interchangeably to refer to the liquid form, sodium hypochlorite, and the solid form, calcium hypochlorite.

Disinfection

- Chlorine compounds are added to water to destroy or inactivate disease-producing (pathogenic) organisms.
- Disinfection differs from sterilization, which is the destruction of all living organisms.

Oxidation

- Hypochlorites are commonly used to oxidize iron, manganese, organic matter, cyanide, and sulfide for subsequent removal.

Taste and Odor Control

- Removal (oxidization) of chemicals and organic matter helps to control tastes and odors in the treated water supply.

Hypochlorite Production

Manufactured Products

- There are two types of hypochlorites that are manufactured: liquid sodium hypochlorite and granular calcium hypochlorite. These will be discussed in more detail later in this unit.

On-Site Generation

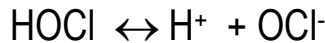
- Hypochlorites can be generated on site. This will be discussed in greater detail in Unit 4 of this module.

ANSI/NSF 60: Water Treatment Chemicals

NSF led the development of American National Standards for all water treatment chemicals. This standard addresses the health effects implementations of treatment chemicals and related impurities. Chemicals used by a water supplier which come in contact with the water may affect the quality of the water. Water treatment chemicals which are certified for conformance with ANSI/NSF Standard 60 are deemed acceptable by the Department. Plant operators should verify that the chemicals used comply with this standard.

Chemistry

- When chlorine is dissolved in water it forms hypochlorous acid.
- Hypochlorous acid (HOCl) dissociates into hydrogen (H⁺) and hypochlorite (OCl⁻) ions.
 - The reaction is almost instantaneous.
 - The reaction is dependent on temperature and pH.
 - The reaction is reversible:



- Together, the hypochlorous acid and hypochlorite ions are considered the “Free Chlorine” available for disinfection. This free chlorine is the most effective disinfectant form of chlorine. We’ll discuss this more in Unit 4.
- Hypochlorous acid is a more effective disinfectant than hypochlorite, so the reaction should be kept in favor of the left side of the equation. This is dependent on pH and temperature.
 - At pH below 6, hypochlorous acid is weak and dissociates poorly.
 - Between pH 6 and 8.5, a sharp change from undissociated HOCl to almost complete dissociation occurs.
 - pH higher than 8.5 hypochlorite predominates, which is not as effective of a disinfectant.
- The normal pH of water supplies is within the range where both hypochlorous acid and hypochlorite ion exist as indicated in the following figure.
 - Solutions of hypochlorites contain excess alkali, which tends to increase pH.

Distribution of HOCl and OCl⁻ in Water

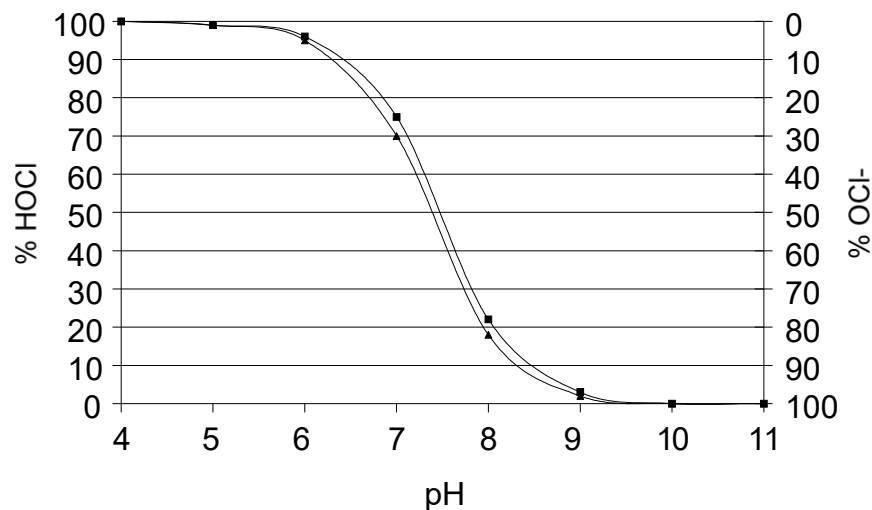


Figure 25.1 The Distribution of HOCl and OCl⁻ in Water

- The disinfection effectiveness of liquid chlorine will decrease with pH increases.

Figure 25.2 shows how the **effectiveness of chlorine decreases as the pH increases**.

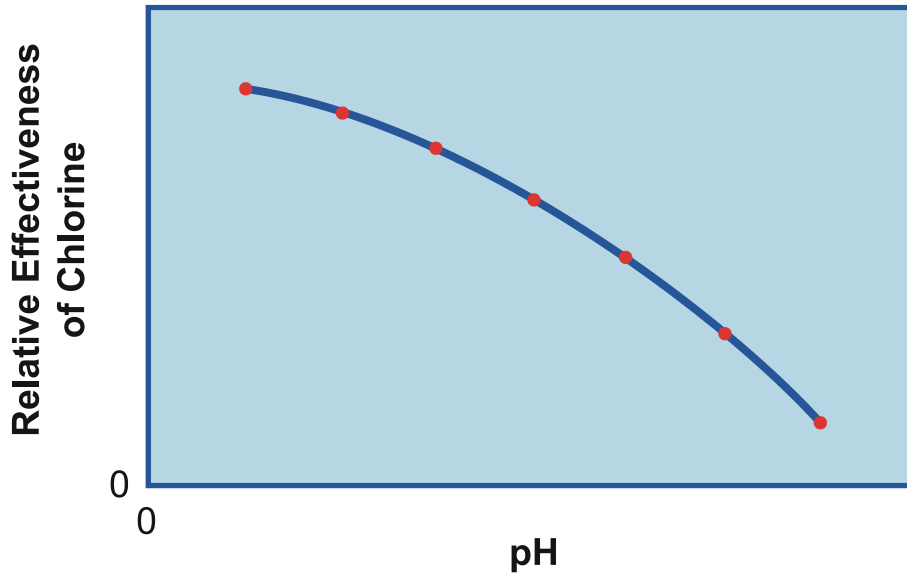


Figure 25.2 The Relationship Between pH and Disinfection When Using Chlorine.

- The disinfection effectiveness of liquid chlorine will increase with temperature.

Figure 25.3 shows how the **effectiveness of chlorine increases as temperature increases**.

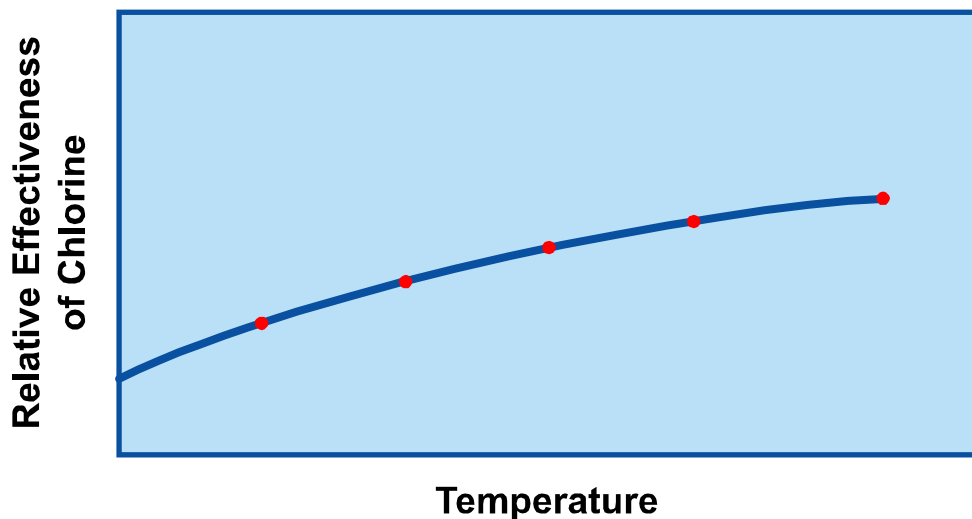


Figure 25.3 The Relationship Between Temperature and Disinfection When Using Chlorine.

Sodium Hypochlorite

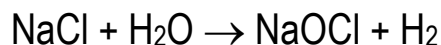
- Sodium hypochlorite is a clear, light yellow-green liquid and is supplied in various strengths.
 - Common household laundry bleach is 5.25% available chlorine by weight.
 - Commercial strength used for water disinfection is 12 – 15% available chlorine by weight. (12.5% solution strength is common).
- It is strongly alkaline and corrosive.
- Sodium hypochlorite has a strong chlorine odor.
- The chemical formula for sodium hypochlorite is NaOCl.
- The properties of sodium hypochlorite are:
 - Boiling point: 110° C (230° F)
 - Specific Gravity: 1.19 (Specific Gravity of Water = 1.0)
 - pH: 9.0 to 12.0
- Reaction in water:



On-Site Generated Liquid Sodium Hypochlorite

A weak solution of liquid sodium hypochlorite can be generated on-site.

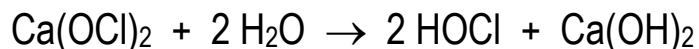
- An electrical charge is applied to a salt brine solution to generate a weak sodium hypochlorite solution. This reaction produces hydrogen gas, which dissipates into the atmosphere.



- The following requirements must be met to generate hypochlorite on-site:
 - Power: 2.5 KWH/lb Cl
 - Salt: 3 lb/lb Cl
 - Water: 15 gal/lb Cl, softened to less than 17 mg/L total hardness
- The properties of on-site generated hypochlorite are:
 - 0.7 to 0.9% available chlorine
 - Solution pH: 9.1 to 9.3

Calcium Hypochlorite - Ca(OCl)₂

- Calcium Hypochlorite has 65 to 70% available chlorine.
 - Every 100 pounds of Ca(OCl)₂ contains 65 – 70 pounds of available chlorine.
 - It is readily soluble in either warm or cold water.
 - Very strong (concentrated) solutions can be prepared. Usually a solution containing 1 to 3 % of available chlorine is prepared. 125 pounds per 1000 gallons of water makes a 1% solution.
- Calcium hypochlorite is a white free-flowing granular powder or solid cake.
 - Its common name is high test hypochlorite (HTH), which is commonly used for swimming pool chlorination.
 - It is stable and can be stored for long periods of time with only a small loss of strength.
- It is readily soluble in either warm or cold water.
 - Very concentrated solutions can be prepared.
 - Solutions are strongly alkaline and corrosive.
- Calcium hypochlorite has a strong chlorine odor.
- The properties of calcium hypochlorite are as follows:
 - Specific Gravity: 2.35 at 20° C (S.G. of Water = 1.0)
 - pH: 11.5 (5% solution)
- Reaction in water:



Stability

Liquid Sodium Hypochlorite

- The stability of hypochlorite solutions is greatly affected by concentration, heat, light, time, and the presence of heavy metals.
 - The higher the concentration, the faster the rate of deterioration.
 - The most stable solutions are of low concentration. On-site generated solution at 1% is more stable than bulk purchased solution at 12%.
 - The higher the temperature, the faster the rate of deterioration.

Table 25.1 The Effects of Temperature of Various Concentrations

Percent Available Chlorine	Half-life (days)		
	140 °F	77 °F	59 °F
10.0	3.5	220	800
5.0	13.0	790	5000
2.5	28.0	1800	--
0.5	100.0	6000	--

- To maintain a somewhat constant solution concentration, quantities should be used within 30 days and stored at 60 to 70 °F.
- Solutions exposed to light deteriorate faster than those kept in darkness. The half-life of 10 to 15 percent solution will be reduced about 3 or 4 times when exposed to light.
- The presence of iron, copper, nickel, or cobalt catalyzes the rate of deterioration; therefore, piping systems should be plastic or other non-metallic material to prevent active corrosion of iron, copper or brass materials.
- **In summary:** The most stable hypochlorite solutions are those with low hypochlorite concentration, a pH of 11 and low iron, copper, or nickel content, stored in the dark at low temperature.

Calcium Hypochlorite

- Dry material is relatively stable under normal atmospheric conditions and will lose 3 to 5 percent available chlorine per year.
 - Reduce deterioration by maintaining a 30 to 60 day stock supply of dry material and mixing fresh solution daily.
- Decomposition is exothermic, meaning it gives off heat. Decomposition proceeds rapidly if heated to 350° F.
 - Many fires of spontaneous origin have been caused by improperly stored calcium hypochlorite.
- Never store calcium hypochlorite where it may be subject to heat or allowed to contact organic material.
 - Calcium hypochlorite releases chlorine fumes when exposed to heat. It may build up pressure in sealed storage containers if containers are exposed to sunlight or other heat sources.
- It should be stored in cool, dry places, isolated from other chemicals since moisture contributes to its deterioration and spontaneous reactions may occur with organic materials and other oxidizable materials.

Vapor Pressure

- All hypochlorite solutions, whether purchased liquid sodium hypochlorite or plant prepared solutions of calcium hypochlorite, will release oxygen gas as the solution decomposes.
- Provisions must be included in the chemical feed system to vent this gas to prevent air binding of the feed lines.
- Care must be taken to prevent trapping the solution between two closed valves since the buildup of pressure may rupture the piping system.



Key points for Unit 1 – Background and Properties

- Chlorine compounds are added to water to destroy or inactivate pathogenic organisms.
- The disinfection effectiveness of liquid chlorine will increase with temperature and decrease as pH increases.
- Hypochlorites can be used to oxidize dissolved metals and to help control taste and odor problems.
- Hypochlorites are provided as liquid sodium hypochlorite or granular calcium hypochlorite.
- Sodium hypochlorite is a clear light yellow-green liquid with a chlorine odor.
- Sodium hypochlorite used for water disinfection is 12 – 15% available chlorine by weight. (12.5% solution strength is common).
- Concentration, heat, light, time, and the presence of dissolved heavy metals can affect the stability of hypochlorite solutions.
- Calcium hypochlorite is a white free-flowing granular powder or a solid cake.
- Calcium hypochlorite is readily soluble in either warm or cold water. However, if it is kept dry and cool, it can be stored for long periods of time.
- Calcium hypochlorite will lose 3 to 5 percent of available chlorine per year. A 30 to 60 day stock supply is often recommended to keep chlorine losses within acceptable limits.
- Calcium Hypochlorite has 65 to 70% available chlorine.



Exercise for Unit 1 – Background and Properties

1. List and explain two uses of hypochlorite.
 - a.

 - b.

2. Matching: Please match the chemical with the available chlorine by weight by drawing lines between:

<u>Chemical</u>	<u>Available Chlorine by Weight</u>
Sodium hypochlorite	0.1 %
Household bleach	5.25%
Calcium hypochlorite	12 to 15%
	65 to 70%

3. Which of the following affect the stability of hypochlorite:
 - Temperature
 - Color
 - Exposure to light
 - How long it is stored

4. Circle the choice that best fills in the blank:
 - The higher / lower the concentration of sodium hypochlorite, the more stable it is.
 - Chlorine is less effective as the temperature decreases / increases .
 - Chlorine is less effective as the pH decreases/ increases.

5. Dry calcium hypochlorite will lose _____ to _____ percent available chlorine per year.

6. All hypochlorite solutions will release oxygen gas as the solution decomposes.
 - a. True _____
 - b. False _____

Unit 2 – Chemical Handling, Storage and Safety

Learning Objectives

- Explain proper handling and storage of hypochlorite
- Use the hypochlorite SDS sheet
- Identify hypochlorite health and environmental hazards
- Identify personal protection equipment and first aid

Quantities

- Provide storage in separate rooms to avoid contact with organic material.
- Storage for a 30-day supply should be available.
- Sodium hypochlorite should not be stored longer than 45 days since its strength decomposes in storage.
- Store in a cool, dry atmosphere.

Types of Storage Containers

Sodium Hypochlorite

- Storage is available in 5, 15, 30 and 55 gallon drums.
- Bulk liquid should be stored in on-site bulk liquid storage tanks.

Calcium Hypochlorite

- Storage is available in 5, 15, 100, and 300 pound cans, and 415 and 800 pound barrels.
- It should be stored in shipping containers until it is used.
- After a container is opened, loss of chlorine will occur, particularly if exposed to moisture in the air.

Storage Rooms

Hypochlorite will spontaneously react with organic materials. As a result, it should be stored separate from all organic materials such as:

- Turpentine.
- Oils.
- Sugar.
- Fats.
- Paper products.
- Other oxidizable materials.

Temperature

- Maintain temperature of storage rooms between 60 and 70° F.

Ventilation

- Forced air ventilation should be provided.
- Operate fans when the room is occupied.

Light

- Prevent exposure to sunlight.

Materials of Construction

Hypochlorites are caustic; therefore, rubber, glass, PVC and other similar materials should be used when contact with hypochlorite solutions is necessary.

Containment

- Provide spill containment for bulk liquid storage facilities.
- Provide containment volume of 110% of the largest liquid vessel to be stored.

Safety Data Sheets

A Safety Data Sheet, or SDS, is available from the chemical manufacturer/supplier for every chemical. The Occupational Safety and Health Administration (OSHA) Hazard Communication Standard of 2012 (HazCom 2012) mandates the use of a single format for safety data sheets featuring 16 sections.

You should read and understand the SDS for each chemical used in the plant. You should also maintain a personal copy for all hazardous chemicals that are used.

An SDS contains detailed assessments of chemical characteristics, hazards and other information relative to health, safety and the environment. The SDS includes:

- **Section 1, Identification**
- **Section 2, Hazard(s) identification**
- **Section 3, Composition/information on ingredients**
- **Section 4, First-aid measures**
- **Section 5, Fire-fighting measures**
- **Section 6, Accidental release measures**
- **Section 7, Handling and storage**
- **Section 8, Exposure controls/personal protection**
- **Section 9, Physical and chemical properties**
- **Section 10, Stability and reactivity**
- **Section 11, Toxicological information**
- **Section 12, Ecological information**
- **Section 13, Disposal considerations**
- **Section 14, Transport information**
- **Section 15, Regulatory information**
- **Section 16, Other information**, includes the date of preparation or last revision.

A sample SDS is located in the Appendix.

Hypochlorite Hazards

- Hypochlorite does not present the same hazards as gaseous chlorine, so it is safer to handle.
- All forms of chlorine can give off chlorine gas and fumes which could irritate one's respiratory system.
 - Chlorine gas is released when hypochlorite is exposed to high heat.
 - Chlorine fumes have a distinct chlorine smell, which can result in respiratory problems similar to those experienced with chlorine gas.
- Hypochlorite should be handled with clean, dry implements that are free of organic materials.

Spills

- Hypochlorite spills should be washed with large volumes of water to dilute it.

Health

- Hypochlorite can irritate both skin and eyes. It will cause discomfort and/or a rash.
- Immediately wash affected areas thoroughly with water, and flush eyes if hypochlorite has come into contact with them.

Environmental

- The effects of hypochlorite on animals are similar to the health effects it has on humans. It is toxic to aquatic life.

Personnel Safety Protection

Basic Equipment

- Safety protection equipment for hypochlorite is similar to equipment used for other corrosive chemicals.
- Equipment must be used and maintained in strict accordance with manufacturers' recommendations and instructions.
- Protective clothing includes:
 - Eye protection.
 - Gloves.
 - Rubber aprons.
- Emergency showers and eye-wash stations should be provided.



Figure 25.4 An Emergency Shower



Figure 25.5 An Eye Wash Station

First Aid

Inhalation

- Remove the injured party and take person(s) to an uncontaminated outdoor area.
- Call for medical assistance.

Skin Contact

- Immediately shower with large quantities of water.
- Remove protective clothing and equipment while in the shower.
- Flush the skin with water for at least 5 minutes.
- Call for medical assistance.
- Keep the affected area cool.

Eye Contact

- Immediately shower with large quantities of water while holding eyes open.
- Call a physician immediately.
- Transfer person(s) promptly to a medical facility.

Ingestion

- Do not induce vomiting.
- Give large quantities of water.
- Call a physician immediately.
- Transfer person(s) promptly to a medical facility.



Key points for Unit 2 – Chemical Handling, Storage and Safety.

- Temperature, ventilation, and light are important considerations when storing hypochlorite.
- Storage rooms should be kept separate and avoid contact with organic material.
- Storage for a 30 day supply is usually recommended.
- Don't store for more than 45 days since hypochlorite decomposes and loses potency.
- Hypochlorites are caustic. Use rubber, glass, PVC and other similar materials when storing hypochlorite.
- Provide spill containment for bulk liquid storage.
- Safety Data Sheets, or SDS, should be readily available at the work site and reviewed by each worker.
- Hypochlorite should be handled with clean, dry implements that are free of organic materials.
- Since hypochlorites are corrosive to the skin, eyes, mouth and respiratory system, workers should be using appropriate protective clothing which includes eye protection, gloves, and rubber aprons. An emergency shower and an eye wash station should also be available.
- Each worker should thoroughly review and/or receive training in first aid procedures for hypochlorite accidents.

**Exercise for Unit 2 – Storage, Handling and Safety**

1. Sodium hypochlorite should not be stored longer than _____ days since its strength decomposes in storage.
2. Calcium hypochlorite should be stored in its _____ containers until it is used.
3. Hypochlorites decompose and release _____ into the air.
4. Forced air ventilation should be turned on whenever workers enter the hypochlorite storage or work area.
 - a. True
 - b. False
5. SDS is an abbreviation for _____
_____.
6. Typical information in a Safety Data Sheet includes:
 - a) The product name and its synonyms.
 - b) Fire and explosion hazard data.
 - c) Toxicity data.
 - d) First aid procedures.
 - e) All of the above.
7. Hypochlorite spills should be washed with large amounts of _____ to dilute it.
8. Hypochlorite will react spontaneously with organic material and should be kept separate from all organic compounds such as: fats, sugar, oils, turpentine, paper, and other oxidizable materials.
 - a) True
 - b) False
9. First aid procedures for skin contact with hypochlorite include showering with large quantities of _____ and calling for medical assistance.
10. Hypochlorite should be stored so that it does not get direct exposure to _____.

References

1 Further information about OSHA requirements can be obtained from:

www.osha.gov

or by phoning the OSHA regional office in Philadelphia 215-861-4900

2 *Safety Practices for Water Utilities*, AWWA Manual of Water Supply Practices, M3, Sixth Edition, 2002.

Unit 3 – Math Principles and Process Control Calculations

Learning Objectives

- Describe math terms, principles and rules for solving equations.
- Review unit cancellation steps.
- Perform calculations for the following types of situations:
 - Calculating changing % concentrations of a chemical
 - Dosage/Feed Rate/Flow
 - Chlorine Demand or Dose
 - CT

Basic Math Principles for Solving Dosage Calculations

Here are some basic math terms and principles.

Fraction: A numerical expression containing a numerator and denominator that represents portions of a whole object. Fractions are used to represent ratios and represent division. For example, the fraction $\frac{1}{4}$ is used to represent the ratio 1:4 and $1 \div 4$.

Parts of a Fraction

1. **Numerator:** The **top** number of a fraction that indicates how many parts are being considered.
2. **Division Line:** The line between the numerator and denominator that indicates that the numerator value is divided by the denominator value to convert a fraction into a decimal.
Example: $1 \div 4 = 0.25$ (as a decimal)
3. **Denominator:** The **bottom** number of a fraction that tells us how many equal parts into which the whole has been divided.

Fraction Written in Vertical Format:

$$\frac{1}{4}$$

(numerator)
(denominator)

Class Activity

Write the following fractions in **vertical** format:

$\frac{1}{8} =$ _____ (The numerator is: _____)

$\frac{2}{6} =$ _____ (The denominator is: _____)

$\frac{8}{10}$ means that _____ is divided by _____

Rules for Solving for an Unknown Variable (such as X)

When solving for the unknown variable (**X**), there are 2 basic objectives:

X must be in the numerator, **AND**

X must be by itself (on one side of the equation).

To accomplish these objectives, **only diagonal movement of terms across the equal sign is permissible in multiplication and division problems.**

$$\frac{(5)}{(3)} = \frac{(3)}{(5)}$$

Explanation of diagonal movement and an example.

An equation is a mathematical statement in which the terms or calculation on one side = the terms or calculation on the other side. To keep both sides equal, any multiplication or division done to one side, must be done to the other. This keeps the equation balanced.

Example 1:

$$5X = 20$$

Question #1 regarding Example #1: Is the **X** in the numerator? _____

Question #2 regarding Example #1: Is the **X** alone on one side of the equation? _____

How do we use diagonal movement to place **X** alone on one side of the equation?

Answer:

Divide both sides by "5" to get **X** alone and **treat both sides of the equation equally.**

Notice that the 5 was moved from the top of the left side to the bottom of the right side of the equation – **a diagonal move.**

$$\frac{5X}{5} = \frac{20}{5}$$

FINAL ANSWER: _____

Example 2:

$$2.5 = \frac{1,000}{X}$$

Question #1 regarding Example #2: Is the **X** in the numerator? _____

How do we move the **X** into the numerator?

Answer:

Multiply both sides of the equation by **X**. Or, you could think of it as simply moving the **X** diagonally from the denominator into the numerator.

$$X(2.5) = \frac{1,000(\cancel{X})}{\cancel{X}} \quad \text{OR} \quad X(2.5) = \frac{1,000}{X}$$

Question #2 regarding Example #2: Is the **X** alone on one side of the equation? _____

How do we use diagonal movement to place **X** alone on one side of the equation?

Answer:

Divide by 2.5 on each side of the equation so that the **X** is alone, but the equation keeps the same value.

$$\frac{X(2.5)}{2.5} = \frac{1,000}{2.5}$$

$$X = \frac{1,000}{2.5}$$

FINAL ANSWER: _____

Class Exercise Solving for X:

$$\frac{X}{200} = 2.4$$

X = _____

$$10 = \frac{3000}{X}$$

X = _____

There are a few rules for doing the various mathematical functions like multiplication, division, addition and subtraction.

Order of Operation for Multiplication, Division, Addition and Subtraction

To solve for **X** when multiplication and division as well as addition and subtraction of terms is indicated, use the following steps:

Simplify as many terms as possible, using the order of operation:

If **brackets or parentheses** contain any arithmetic, **simplify within these groups first** by:

Completing the multiplication or division, **THEN**

Complete the addition or subtraction.

Complete all **multiplication and division** from left to right, **THEN**

Complete all **addition and subtraction** from left to right.

Verify that the **X** term is in the numerator. If it is not, move the **X** term to the numerator, using a diagonal move.

Verify that **X** is by itself, on one side of the equation.

Explanation of the Order of Operation and an example.

Example 1:

$$4\left(\frac{10}{2} + 3\right) + 6$$

Step #1: Simplify terms **within parentheses** by **multiplying and dividing** from left to right.

$$= 4(5 + 3) + 6$$

Step # 2: THEN simplify terms **within parentheses** by **adding and subtracting** from left to right.

$$= 4(8) + 6$$

Step #3: Simplify terms **outside of parentheses** by **multiplying and dividing** from left to right.

$$= 32 + 6$$

Step #4: Simplify terms **outside of parentheses** by **adding and subtracting** from left to right.

$$= 38$$

Class Exercise: Explanation of the Order of Operation

Example 2:

$$(X)(2)(8.34) = 500$$

Step #1: Simplify terms by **multiplying and dividing** from left to right.

Step # 2: Verify that the X term is in the numerator. If it is not, move the **X** term to the numerator, using a diagonal move.

Is X in the numerator _____

Step #3: Verify that **X** is by itself, on one side of the equation. If it is not, divide both sides of the equation by the number on the **X** side of the equation.

Does this equation require a division? _____ If so by what number? _____

FINAL ANSWER: $500 \div 16.68 =$ _____

Now let's discuss the concept of unit cancellation.

Problem Solving Using Unit Cancellation:

We give it that name because you cancel units until the problem is solved.

- Unit cancellation involves canceling units in the numerator and denominator of unit fractions to obtain the desired units of measurement.
- Unit cancellation can be used to make conversions or to solve problems.

There are three basic rules for using unit cancellation on the next page.

Basic rules for using unit cancellation:

Unit fractions should be written in a vertical format. A unit fraction has one unit in the numerator (above the line) and one unit in the denominator (below the line).

1. A fraction is structured like this:
$$\frac{\text{numerator}}{\text{denominator}}$$

For example, gallons per minute (GPM) should be written as
$$\frac{\text{gal}}{\text{min}}$$

2. Any unit which appears in the numerator of one unit fraction and the denominator of another unit fraction is canceled.

The following is an example of how units are canceled:

$$20 \frac{\text{gal}}{\text{min}} \times 60 \frac{\text{min}}{\text{hr}} = 1200 \frac{\text{gal}}{\text{hr}}$$

3. It may be necessary to invert data and the corresponding units.

$$10 \frac{\text{gal}}{\text{min}} \text{ is the same as } \frac{1 \text{ min}}{10 \text{ gal}}$$

Example Problem: How many mL/min are in a flow of 5 gal/day?

Problem Set Up

List all known and unknown data.

Unknown: $\frac{? \text{ mL}}{\text{min}}$

Known: $\frac{5 \text{ gal}}{\text{day}}$

From this data, find units which need to be converted.

Volume units: gal to mL

Time units: day to min

Find conversions from the conversion charts.

$$1 \text{ gal} = 3,785 \text{ mL}$$

$$1 \text{ day} = 1,440 \text{ min}$$

This is all the data we need to use in the problem.

Steps to solving problems using unit cancellation

Step 1: List unknown data including units in vertical format followed by an equal sign.

Example: Unknown data: $\frac{? \text{ mL}}{\text{min}} =$

NOTE: You may need to invert data throughout the following steps.

Step 2: Find data (known or a conversion) that has the same numerator unit as the unknown numerator. Place it to the right of the equal sign. Add a multiplication sign. This positions your numerator.

Example: $\frac{? \text{ mL}}{\text{min}} = 3,785 \frac{\text{mL}}{1 \text{ gal}} \times$ *Conversion*

Step 3: To cancel unwanted denominator unit, find data (known or a conversion) that has the same numerator unit. Place it to the right of data used in Step 2. Place a multiplication sign between each piece of data.

Example: $\frac{? \text{ mL}}{\text{min}} = 3,785 \frac{\text{mL}}{1 \text{ gal}} \times \frac{5 \text{ gal}}{1 \text{ day}} \times$ *Known*

Step 4: Continue to place data (known or a conversion) into equation to systematically cancel all unwanted units until only the unknown denominator units remain.

Example: $\frac{? \text{ mL}}{\text{min}} = 3,785 \frac{\text{mL}}{\text{gal}} \times \frac{5 \text{ gal}}{1 \text{ day}} \times \frac{1 \text{ day}}{1,440 \text{ min}}$ *Conversion*

Note 1: All units must cancel, leaving only the units you are solving for in the unknown data. If all units except the unknown units are not crossed out, check the list of known data to see if all relevant known data was used to solve the problem and all necessary conversions were made.

Note 2: If you need to invert the known data or conversion values and units to cancel, remember to move the appropriate units with the value.

Step 5: Multiply the values of all numerators and place this value in the numerator of the answer. Multiply the values of all denominators and place this value in the denominator of the answer. Divide to calculate the final answer.

Important: Check the answer to verify that the value is reasonable.

Example: $\frac{? \text{ mL}}{\text{min}} = \frac{18,925 \text{ mL}}{1440 \text{ min}} = 13.1 \frac{\text{mL}}{\text{min}}$

Practice Problem:

How many hours will it take to empty a 55 gallon drum of sodium hypochlorite using a chemical feed pump that's pumping at a rate of 30 mL/min?

Solution:

Unknown Data: _____

Known Data: _____

Unit Cancellation Steps:

If a water treatment plant produces 3 million gallons per day, and uses chlorine gas, dosed at 7 mg/l, you can calculate how many pounds per day the plant will use.

Unit Cancellation Steps:

Positions the numerator

Known Known

$$? \frac{\text{lbs}}{\text{day}} = \frac{1 \text{ lb}}{454 \text{ g}} \times \frac{1 \text{ g}}{1000 \text{ mg}} \times \frac{7 \text{ mg}}{\text{L}} \times \frac{3.785 \text{ L}}{1 \text{ gal}} \times \frac{1,000,000 \text{ gal}}{1 \text{ MG}} \times \frac{3 \text{ MG}}{1 \text{ day}} = \frac{79485000 \text{ lb}}{454000 \text{ day}} = \frac{175 \text{ lbs}}{\text{day}}$$

Conversions Conversions

Cancel unwanted units until you get to the denominator unit that you are solving the equation for (in this example it's "day")

Unit Cancellation Steps

Step 1: List ? unknown data including units followed by an = sign

Step 2: Place data with same numerator unit to the right of the equal sign followed by a multiplication sign. This positions your numerator.

Step 3: To cancel unwanted denominator unit, next place data with same numerator unit.

Step 4: Continue to place data into equation to systemically cancel all unwanted units until only the unknown denominator units remain.

Step 5: Do the math (Multiply all numerator values, multiply all denominator values, then divide numerator by the denominator.)

Example:

The density of a liquid is 1 gm/mL in the Metric system. **What is the density in the English system (lbs/gal)?**

Notice that this conversion proves that the density of water = 8.34 lbs/gallon.

$$? \frac{\text{lbs}}{\text{gal}} = \frac{1 \text{ lb}}{454 \text{ gm}} \times \frac{1 \text{ gm}}{1 \text{ mL}} \times \frac{3785 \text{ mL}}{1 \text{ gal}} = 8.34 \frac{\text{lbs}}{\text{gal}}$$

Helpful Hints:

Numerator
Denominator

Vertical format: $5 \text{ gal} = \frac{5 \text{ gal}}{1}$

1 gm = 1000 mg is written: $\frac{1 \text{ gm}}{1000 \text{ mg}}$ OR $\frac{1000 \text{ mg}}{1 \text{ gm}}$

“per” means divided by: Ex. $5 \text{ gpm} = \frac{5 \text{ gal}}{1 \text{ min}}$

Inverting: $\frac{5 \text{ gal}}{1 \text{ min}} = \frac{1 \text{ min}}{5 \text{ gal}}$

Calculations for Changing % Concentrations of a Chemical

We purchase sodium hypo in 12%-15% strength, but we may have to dilute it (maybe our pump is just sized too big). Then we have to figure out how many gallons we are going to pump.

Problem: In 12 hours, you feed 1.2 gallons of 12% hypochlorite solution. How many gallons would you have to use if the concentration is 5%?

Step 1: Set up math equation: Remember, in an equation, the terms or calculations on one side equal the terms or calculations on the other side. We can use this fact to put our problem in equation form.

$$1.2 \text{ gal} \times 12\% = ? \text{ gal} \times 5\%$$

Step 2: Divide both sides by 5% to get ? gal alone on the right side of the equation (basic math rule)

$$\frac{1.2 \text{ gal} \times 12\%}{5\%} = ? \text{ gal} \times \frac{5\%}{5\%}$$

Step 3: Multiply $1.2 \times 12 = 14.4$ in the numerator (basic math rule)

Step 4: Perform the division: 14.4 (numerator) \div 5 (denominator) = 2.88 gallons (final basic math rule)

$$\frac{14.4}{5}$$

Note – the 12 hours was not important!

CALCULATIONS: CHANGING % CONCENTRATIONS

Practice Problem: You purchase a new pump. The old pump fed 5.5 gallons daily of 15% sodium hypochlorite. You need to change your concentration to a 6% solution. How many gallons can you now expect to use each day?

Step 1: Set up math equation: _____ gal x _____% = ? gal x _____%

Step 2: Divide both sides by 6% to get ? gal alone on the right side of the equation

$$\frac{\text{_____ gal}}{6\%} \times \text{_____}\% = ? \text{ gal} \times \frac{6\%}{6\%}$$

Step 3: Multiply _____ x _____ = _____ in the numerator

Step 4: Perform the division: _____ (numerator) ÷ _____ (denominator) = _____ gallons

Process Calculations

There are three basic chlorination process calculations:

1. chlorine dosage or feed rate
2. chlorine demand
3. CT calculation

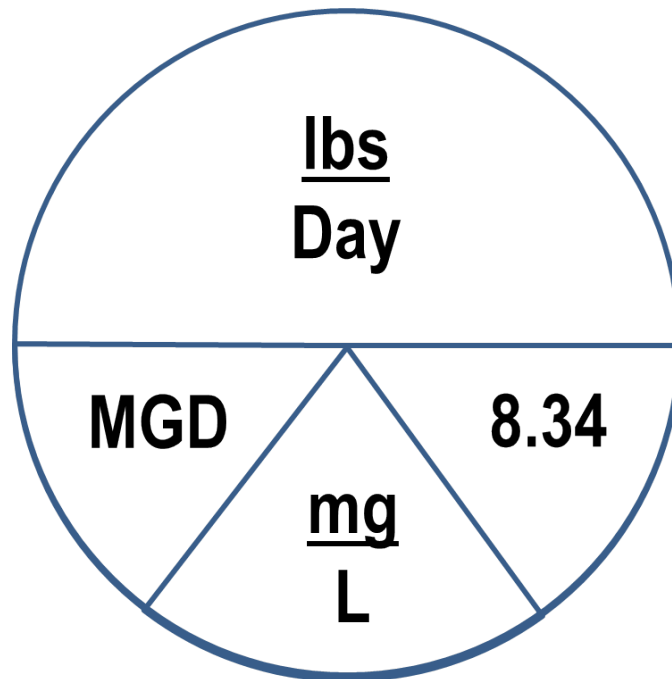
Chlorine Dosage/Feed Rate Calculation

- To perform the calculation, you will need to know the amount of chlorine being added and the amount of water being treated.

$$\text{Feed Rate, lbs/day} = \text{Flow (MGD)} \times \text{Dosage (mg/L)} \times 8.34 \text{ lbs/gal}$$

This formula is represented in the following diagram called the Davidson Pie which was created by Gerald Davidson, Manager, Clear Lake Oaks Water District, Clear Lake Oaks, CA.

Davidson Pie



Key Acronyms:

MG = million gallons

MGD = million gallons per day

Davidson Pie Diagram Interpretation and Formulas

This diagram can be used to solve for 3 different results: dosage, feed rate, and flow (or volume).

As long as you have 2 of those 3 variables, you can solve for the missing variable.

Davidson Pie InterpretationMiddle line = divided by (\div)Bottom diagonal lines = multiply by (\times)

In other words, here are the 3 equations that can be used with these variables:

1. **Feed Rate, lbs/day = Flow (MGD) or Volume (MG) \times Dosage (mg/L) \times 8.34** (which is the density of water)
2. **Flow (MGD) = lbs/day \div (Dosage, mg/L \times 8.34)**

$$\text{Vertical Format: Flow(MGD) = } \frac{\text{Feed Rate (lbs/day)}}{[\text{Dosage (mg/L) } \times \text{ 8.34}]}$$

3. **Dosage (mg/L) = lbs/day \div (Flow, MGD \times 8.34)**

$$\text{Vertical Format: Dosage (mg/L) = } \frac{\text{Feed Rate (lbs/day)}}{[\text{Flow(MGD) } \times \text{ 8.34}]}$$

Equation # 1: Feed Rate Calculation Using Flow

Solving for Pounds/Day

Feed Rate, Pounds per day = flow(MGD) x dose(mg/L) x (8.34)

This equation alone (without extra steps) can be used to solve for feed rates of 100% strength chemicals, such as chlorine gas.

A water treatment plant produces 3 million gallons per day, and uses chlorine gas, dosed at 7 mg/L, how many pounds per day will the plant use?

Feed Rate, Pounds per day = $3 \times 7 \times 8.34 = 175$ pounds per day (same answer as the unit cancellation answer from earlier)

When we are using chemicals that aren't 100% strength, we need to factor in the active strength of the chemical into the calculation.



Active Strength is the percentage of a chemical or substance in a mixture that can be used in a chemical reaction. (also referred to as % purity)

Feed Rate Calculations Using Flow with a % Strength (i.e., % pure) Solution

Unlike chlorine gas, sodium and calcium hypochlorite solutions are not 100 percent pure. For example, the sodium hypochlorite typically used is 12.5% pure. That means that out of every gallon of hypochlorite, only 12.5% is the chlorine component, and the other material (87.5%) is not chlorine.

Problem: A water plant uses sodium hypochlorite (12.5%) to disinfect the water. The target dose is 1.2 mg/L. They treat 0.25 million gallons per day. How many pounds of sodium hypochlorite will need to be fed?

Step 1: Solve for pounds per day (feed rate) for 100% pure chemical (no impurities).

Using the formula pounds per day = flow x dose x 8.34 = $(0.25)(1.2)(8.34) = 2.5$ pounds of chlorine is required.

CALCULATIONS: DOSAGE/FEED RATE/FLOW

Step 2: Calculate # of pounds of 12.5% solution needed to achieve Step 1 feed rate.

Since they are using hypochlorite, and only 12.5% of the hypo is chlorine, we need to calculate how many pounds of hypo are required to get 2.5 pounds of chlorine. To do that we need to change the percent to a decimal, and divide that into the pounds required.

a) Convert % purity of solution into a decimal:

$$\frac{12.5\%}{100\%} = 0.125$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{2.5 \text{ pounds}}{0.125 (\% \text{ purity as a decimal})} = 20 \text{ pounds of 12.5\% hypochlorite.}$$

TIP: Answer will always be more pounds than Step 1 result because solution is not 100% pure.

Practice Problem: A water plant uses sodium hypochlorite (15%) to disinfect the water. The target dose is 1.6 mg/L. They treat 0.25 million gallons per day. How many pounds of sodium hypochlorite will need to be fed?

Step 1: Solve for pounds per day (feed rate) for 100% pure chemical (no impurities).

Using the formula pounds per day = flow x dose x 8.34 = (____)(____)(____) = _____ pounds of chlorine is required.

Step 2: Calculate # of pounds of 15% solution needed to achieve Step 1 feed rate.

a) Convert % purity of solution into a decimal:

$$\frac{15\%}{100\%} = \underline{\hspace{2cm}}$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{\underline{\hspace{2cm}} \text{ pounds}}{0.15 (\% \text{ purity as a decimal})} = \underline{\hspace{2cm}} \text{ of 15\% hypochlorite.}$$

TIP: Answer will always be more pounds than Step 1 result because solution is not 100% pure.

When you have a flow in gallons per minute (GPM) or gallons per day (GPD), you will need to convert those values into million gallons per day (MGD) before using the feed rate formula.

Converting from GPM to MGD before solving with the formula

Problem: A water treatment plant operates at the rate of 75 gallons per minute. They dose soda ash at 14 mg/L. How many pounds of soda ash will they use in a day?

We'll use unit cancellation to show the steps involved in doing this math.

Step 1: Convert gallons per minute into gallons per day using unit cancellation.

$$? \frac{\text{gal}}{\text{day}} = 75 \frac{\text{gal}}{\text{minute}} \times \frac{60 \text{ minutes}}{\text{hour}} \times \frac{24 \text{ hours}}{\text{day}} = 108,000 \frac{\text{gallons}}{\text{day}}$$

Step 2: Convert gallons per day into million gallons per day (MGD) using unit cancellation.

$$? \frac{\text{MG}}{\text{day}} = 1 \frac{\text{MG}}{1,000,000 \text{ gallons}} \times \frac{108,000 \text{ gallons}}{\text{day}} = 0.108 \text{ MGD}$$

Step 3: Use MGD in feed rate formula to solve for lbs/day

Feed Rate, Pounds per day = flow(MGD) x dose(mg/L) x (8.34)

$$0.108 \times 14 \times 8.34 = 12.61 \text{ lbs/day}$$

Steps 1 and 2 can be combined like this:

Step 1: Convert gallons per minute into million gallons per day (MGD) using unit cancellation.

$$? \frac{\text{MG}}{\text{day}} = 1 \frac{\text{MG}}{1,000,000 \text{ gallons}} \times 75 \frac{\text{gal}}{\text{minute}} \times \frac{60 \text{ minutes}}{\text{hour}} \times \frac{24 \text{ hours}}{\text{day}} = 0.108 \frac{\text{MG}}{\text{day}}$$

You can also use the conversion of 1 day = 1440 mins (60 X 24) to remove 2 conversions (minutes/hour and hours/day)

Step 1: Convert gallons per minute into million gallons per day (MGD) using unit cancellation.

$$? \frac{\text{MG}}{\text{day}} = 1 \frac{\text{MG}}{1,000,000 \text{ gallons}} \times 75 \frac{\text{gal}}{\text{minute}} \times 1440 \frac{\text{minutes}}{\text{day}} = 0.108 \frac{\text{MG}}{\text{day}}$$

Converting from GPD to MGD before solving with the formula

Problem: If a water treatment plant is making water at the rate of 150,000 gallons per day, and the chlorine dose is 0.8 mg/L, how many pounds of chlorine will they use daily (assume 100% strength)?

Step 1: Convert gallons per day into million gallons per day (MGD) using unit cancellation.

$$? \frac{\text{MG}}{\text{day}} = 1 \frac{\text{MG}}{1,000,000 \text{ gallons}} \times \frac{150,000 \text{ gallons}}{\text{day}} = 0.15 \text{ MGD}$$

Step 2: Use MGD in feed rate formula to solve for lbs/day

Feed Rate, Pounds per day = flow(MGD) x dose(mg/L) x (8.34)

$$0.15 \times 0.8 \times 8.34 = 1 \text{ lb/day}$$

Notice that you can also use this first equation and substitute **volume** for flow. In addition, the feed rate is determined in **pounds**, rather than pounds per day.

Feed Rate Problem Using Volume instead of Flow with a % Strength Chemical

Your storage tank has been taken out of service for cleaning. The 50,000 gallon tank must be properly disinfected before you can return it to service. The consulting firm recommends you use 25 mg/L of 68% calcium hypochlorite. How many pounds of calcium hypochlorite do you need to add to the water?

Step 1: Convert volume (in gallons) into MG so that the feed rate formula can be used.

$$? \text{MG} = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times 50,000 \text{ gal} = 0.05 \text{ MG}$$

Step 2: Solve for pounds (feed rate) for 100% pure chemical (no impurities).

$$? \text{ lbs} = \text{volume(MG)} \times \text{dose(mg/L)} \times 8.34 = (0.05)(25)(8.34) = 10.4 \text{ pounds of chlorine is required.}$$

Step 3: Calculate # of pounds of 68% solution needed to achieve Step 2 feed rate.

Since they are using hypochlorite, and only 68% of the hypo is chlorine, we need to calculate how many pounds of hypo are required to get 10.4 pounds of chlorine. To do that we need to change the percent to a decimal, and divide the pounds required by the purity of the solution (as a decimal).

a) Convert % purity of solution into a decimal:

$$\frac{68\%}{100\%} = 0.68$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{10.4 \text{ pounds}}{0.68 \text{ (\% purity as a decimal)}} = 15.3 \text{ pounds of hypochlorite.}$$

Practice Problem #1

Calculate the amount of calcium hypochlorite required to dose a 500,000 gallon storage tank to a dose of 25 mg/L using granular calcium hypochlorite that indicates it is 65% chlorine.

Step 1: Convert volume (in gallons) into MG so that the feed rate formula can be used.

$$? \text{ MG} = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times (\text{_____}) \text{ gal} = \text{_____ MG}$$

Step 2: Solve for pounds (feed rate) for 100% pure chemical (no impurities).

$$? \text{ lbs} = \text{volume(MG)} \times \text{dose(mg/L)} \times 8.34 = (\text{_____})(\text{_____})(\text{_____}) = \text{_____ pounds of chlorine is required.}$$

Step 3: Calculate # of pounds of 65% solution needed to achieve Step 2 feed rate.

a) Convert % purity of solution into a decimal:

$$\frac{65\%}{100\%} = \text{_____}$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{\text{_____ pounds}}{0.65} = \text{_____ pounds of 65\% calcium hypochlorite.}$$

Practice Problem #2:

Calculate the amount of chlorine required for a dosage of 1 mg/L in a 600,000 gallon storage tank. The tank is $\frac{3}{4}$ full. (Assume 100% strength)

Step 1: Calculate volume of tank that is not 100% full by multiplying the volume by the fraction (or its equivalent decimal.)

$$600,000 \times \frac{3}{4} = \underline{\hspace{2cm}} \text{ gallons} \quad \text{OR} \quad 600,000 \times 0.75 = \underline{\hspace{2cm}} \text{ gallons}$$

Step 2: Convert volume (in gallons) into MG so that the feed rate (lbs) formula can be used.

$$? \text{ MG} = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times 450,000 \text{ gal} = \underline{\hspace{2cm}} \text{ MG}$$

Step 3: Solve for pounds (feed rate) for 100 % pure chemical (no impurities).

$$? \text{ lbs} = \text{volume(MG)} \times \text{dose(mg/L)} \times 8.34 = (.45)(1)(8.34) = \underline{\hspace{2cm}} \text{ pounds of chlorine is required.}$$

The following page provides summary of the feed rate calculation steps we use when we have either a volume or a flow. **This will always be used when solving for POUNDS or POUNDS/DAY. So, whenever you see a problem that asks to solve for an amount in pounds, use these steps!**

**Summary of Steps for Solving Feed Formula Calculations in Pounds/Day
for % Strength (i.e., % Purity) Solutions
(using either volume or flow rate)**

Example: Calculate the amount of calcium hypochlorite required to dose a 500,000 gallon storage tank to a dose of 25 mg/L using granular calcium hypochlorite that indicates it is 65% chlorine.

Step 1: Convert volume (in gallons) into MG or flow in gallons (per day or per minute) into MGD so that the feed rate (lbs or lbs/day) formula can be used.

$$?MG = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times \text{volume of tank (gal)} \text{ OR}$$

$$?MGD = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times \frac{\text{volume of flow (gal)}}{1 \text{ day}} \text{ OR}$$

$$?MGD = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times \frac{\text{volume of flow (gal)}}{1 \text{ min}} \times \frac{1440 \text{ min}}{\text{day}}$$

Step 2: Solve for pounds (feed rate) for 100% pure chemical (no impurities). (104.25 pounds for this example)

$$? \text{ lbs} = \text{volume(MG)} \times \text{dose(mg/L)} \times 8.34 = \text{pounds of chlorine that are required.}$$

Step 3: Calculate # of pounds of % solution needed (in this example, 65%) to achieve Step 2 feed rate.

a) Convert % purity of solution into a decimal:

$$\frac{65\%}{100\%} = 0.65$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{104.25 \text{ pounds}}{0.65} = 160.39 \text{ pounds of 65\% calcium hypochlorite.}$$

TIP: Answer will always be more pounds than Step 2 result because solution is not 100% pure.

How do we convert pounds of a solution into gallons of that same solution?

Solving for Gallons/Day

Calculating “Active Ingredient” Weight

In addition to knowing that solutions are not 100% pure (i.e. 100% active), we also need to determine the weight of the active strength ingredients within that solution.



Active ingredient weight is the number of pounds of “active ingredient” per gallon of a % solution that cause a chemical reaction.

- This “active ingredient” weight value is then used in a calculation with the 100% pure “lbs/day” feed rate to determine the “gal/day” feed rate.

We need the specific gravity of the solution which is found on the SDS sheet to calculate the weight of the active ingredients within that solution.

Calculating the Weight of the “Active Ingredient” of a % Solution Chemical

EXAMPLE: How many pounds of chlorine are there in a gallon of sodium hypochlorite that is 12.5% pure that has a specific gravity of 1.15?

Step 1: Solve weight equation (lbs/gal) for 1 gallon of chemical

Weight, lbs/gal = (Specific gravity of substance) x (weight of a gallon of water)

$$1.15 \quad \times \quad 8.34 \frac{\text{pounds}}{\text{gallon}} = 9.59 \frac{\text{pounds}}{\text{gallon}}$$

Step 2: Determine the “active ingredient” weight of the chlorine based on the % purity of solution

a) Convert % purity of solution into a decimal:

$$\frac{12.5\%}{100\%} = 0.125$$

b) Multiply the weight of a gallon (from step 1) by the % purity of the product (as a decimal).

$$9.59 \frac{\text{pounds}}{\text{gallon}} \times 0.125 = 1.2 \text{ pounds of available chlorine in a gallon of 12.5\% sodium hypochlorite}$$

This “active ingredient” weight provides the pounds of available chlorine that is found in each gallon of 12.5% sodium hypochlorite solution. Within the 9.59 pounds of 12.5% sodium hypochlorite, there are 1.2 pounds of available chlorine (i.e., active ingredient).

Now that we know the active ingredient weight of the solution, we can use this information to determine how many gallons we need to feed. We'll use unit cancellation to determine the number of gallons/day from the lbs/day feed rate.

Using "Active Ingredient" Weight to Convert Feed Rate from lbs/day to gals/day

Example: A water plant uses sodium hypochlorite (12%) to disinfect the water which provides 1.2 lbs/gal of available chlorine ("active ingredient" weight). The required dosage is 2.5 mg/L. They treat 118,000 gallons per day. How many **gallons** of sodium hypochlorite will need to be fed?

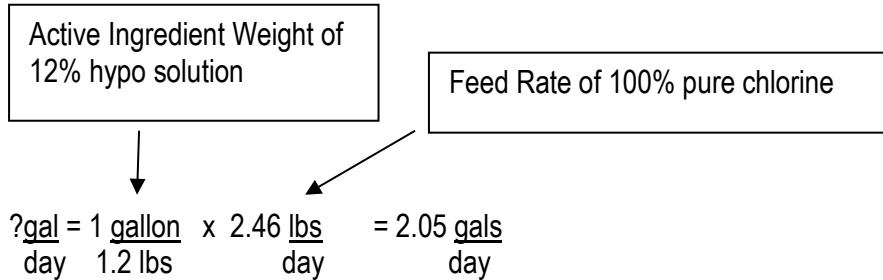
Step 1: Convert flow in gallons (per day) into MGD so that the feed rate (lbs/day) formula can be used.

$$? \text{ MGD} = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times \frac{118,000 \text{ (gal)}}{1 \text{ day}} = 0.118 \text{ MGD}$$

Step 2: Solve for pounds per day (feed rate) for 100% pure chemical (no impurities).

Using the formula pounds per day = flow x dose x 8.34 = (0.118)(2.5)(8.34) = 2.46 pounds of chlorine is required.

Step 3: Use "active ingredient" weight with unit cancellation steps to convert lbs/day to **gals/day**



NOTE: When you are given the "active ingredient" weight of a solution to solve a feed rate problem, you do not need to use the % purity factor because it was used in the weight calculation.

Practice Problem: A water plant uses sodium hypochlorite (12.5%) to disinfect the water which provides 1.2 lbs/gal of available chlorine (“active ingredient” weight). The chlorine dosage is 1.6 mg/L. They treat 600,000 gallons per day. How many gallons of sodium hypochlorite will need to be fed?

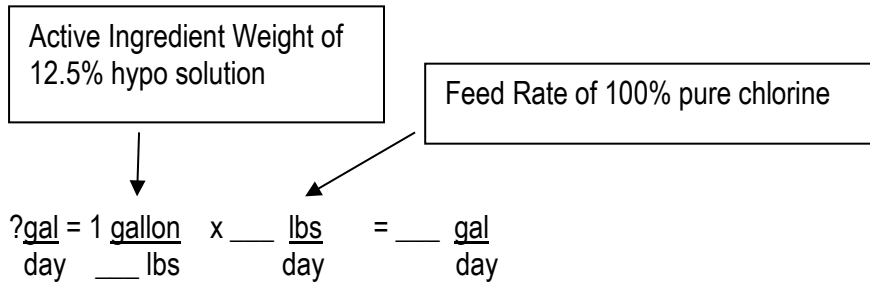
Step 1: Convert flow in gallons (per day) into MGD so that the feed rate (lbs/day) formula can be used.

$$? \text{ MGD} = \frac{1 \text{ MG}}{1,000,000 \text{ gal}} \times \frac{600,000 \text{ gal}}{1 \text{ day}} = \text{_____ MGD}$$

Step 2: Solve for pounds per day (feed rate) for 100% pure chemical (no impurities).

Using the formula pounds per day = flow x dose x 8.34 = (____)(____)(8.34) = ____ pounds of chlorine is required.

Step 3: Use “active ingredient” weight with unit cancellation steps to convert lbs/day to gals/day



NOTE: When you are given the “active ingredient” weight of a solution to solve a feed rate problem, you do not need to use the % purity factor because it was used in the weight calculation.

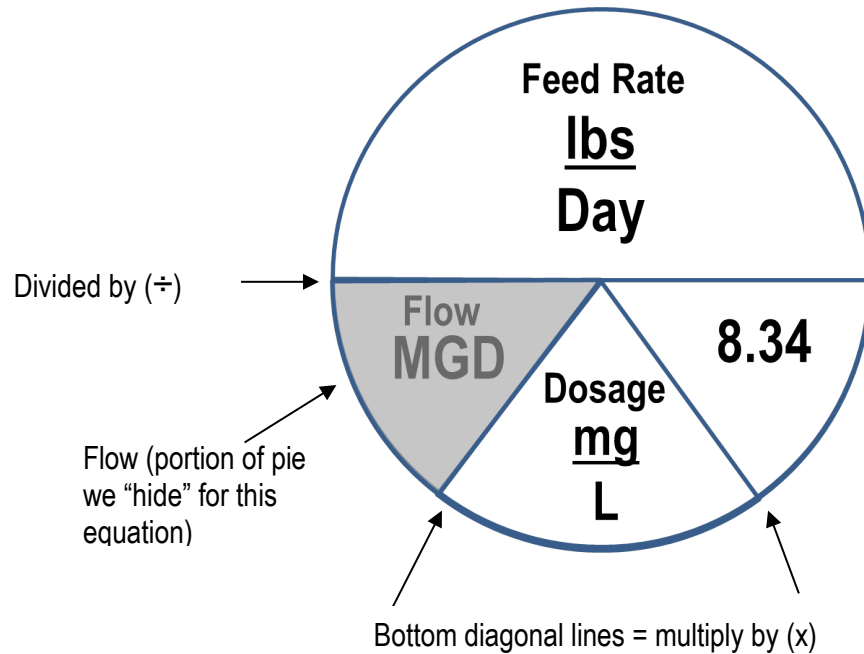
Tip: Two Different Paths for Solving

If the problem **asks for pounds or pounds per day**, solve the equation using the feed rate formula and dividing by the percent solution. We’ll call this **Path 1** (steps are summarized on page 3-23)

If the problem **asks for gallons or gallons per day**, solve the equation using the feed rate equation and available chlorine (or active ingredient weight). We’ll call this **Path 2** (steps summarized are shown on page 3-25)

Equation #2: Solving for Flow Using the Feed Rate Formula

We will continue to use the Davidson Pie Feed Rate/Dosage Formula.



Remember: Middle line represents a division sign (÷)
 Bottom diagonal lines = multiply by (x)

Vertical Format: **Flow, MGD** =
$$\frac{\text{Feed Rate, lbs/per day}}{(\text{Dose, mg/L})(8.34)}$$

Problem: A water treatment plant uses 14 pounds of chlorine to treat their water daily. The chlorine dose is 1.5 mg/L. What is their flow rate in MGD?

Step 1: Set up the variables in vertical format and insert known values

$$? \text{ Flow (MGD)} = \frac{\text{Feed Rate, lbs/day}}{(\text{Dose})(8.34)} = \frac{14 \text{ lbs/day}}{(1.5)(8.34)}$$

← Known feed rate
↑
Known Dose

Step 2: Multiply 1.5 x 8.34 in the denominator = 12.51 (basic math rule)

Step 3: Perform the **FLOW** division: 14 (numerator) ÷ 12.51 (denominator) = 1.1 MGD

Practice Problem: A water treatment plant uses 8 pounds of chlorine daily and the dose is 17 mg/L. How many gallons are they producing?

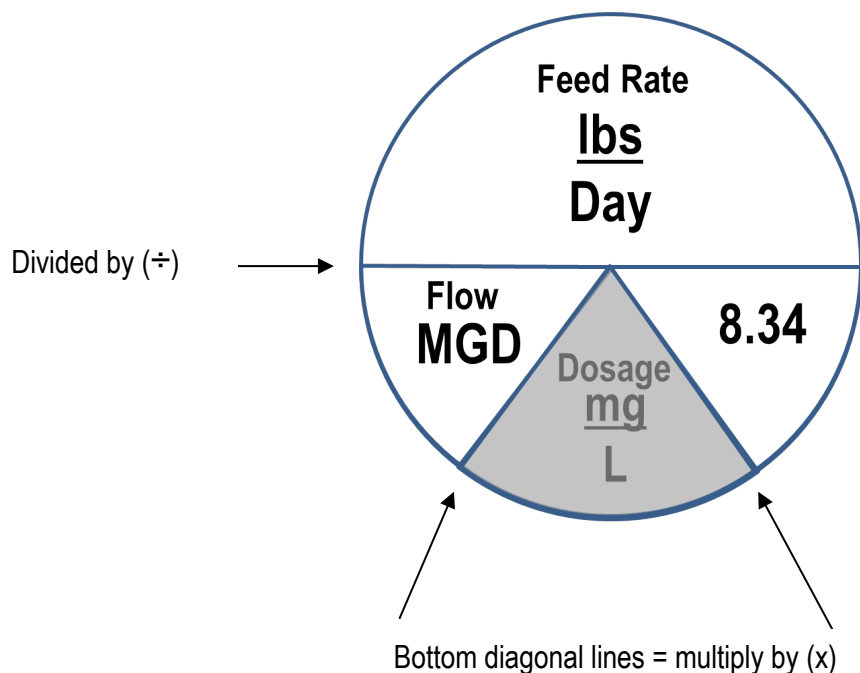
Step 1: Set up the variables in vertical format and insert known values

$$? \text{ Flow (MGD)} = \frac{\text{Feed Rate, lbs/day}}{(\text{Dose})(8.34)} = \frac{(\quad) \text{ lbs/day}}{(\quad)(8.34)}$$

Step 2: Multiply $__ \times 8.34$ in the denominator = $______$ (basic math rule)

Step 3: Perform the **FLOW** division: 8 (numerator) \div 141.78 (denominator) = $______ \text{ MGD}$

Equation #3: Solving for Dose Using the Feed Rate Formula



Vertical Format: $\text{Dose, mg/L} = \frac{\text{Feed Rate, lbs/per day}}{(\text{Flow, MGD})(8.34)}$

Example: A water treatment plant is producing 1.5 million gallons per day of potable water, and uses 38 pounds of chlorine each day. What is the chlorine dose at that plant?

Step 1: Set up the variables in vertical format and insert known values

$$? \text{ Dose (mg/L)} = \frac{\text{Feed Rate, lbs/day}}{(\text{Flow, MGD})(8.34)} = \frac{38 \text{ lbs/day}}{(1.5)(8.34)}$$

← Known feed rate
↑
Known Flow (MGD)

Step 2: Multiply 1.5 x 8.34 in the denominator = 12.51 (basic math rule)

Step 3: Perform the **DOSE** division: 38 (numerator) ÷ 12.51 (denominator) = 3.03 mg/L

Practice Problem: A water treatment plant produces 150,000 gallons of water every day. It uses an average of 2 pounds of permanganate for iron and manganese removal. What is the dose of the permanganate?

Step 1: Set up the variables in vertical format.

$$? \text{ Dose (mg/L)} = \frac{\text{Feed Rate, lbs/day}}{(\text{Flow, MGD})(8.34)}$$

Step 2: Convert gallons per day into MGD and insert known values into equation.

$$? \frac{\text{MG}}{\text{day}} = 1 \frac{\text{MG}}{1,000,000 \text{ gallons}} \times \frac{150,000 \text{ gallons}}{\text{day}} = \underline{\hspace{1cm}} \text{ MGD}$$

$$? \text{ Dose (mg/L)} = \frac{\text{Feed Rate, lbs/day}}{(\text{Flow, MGD})(8.34)} = \frac{(2) \text{ lbs/day}}{(\underline{\hspace{1cm}})(8.34)}$$

Step 3: Multiply x 8.34 in the denominator = (basic math rule)

Step 4: Perform the **DOSE** division: 2 (numerator) ÷ 1.25 (denominator) = $\frac{\text{mg}}{\text{L}}$

Chlorine Demand or Dose Calculation

- A sufficient amount of chlorine must be added so that the chlorine demand is met and the desired chlorine residual is provided.

$$\text{Chlorine Demand (mg/l)} = \text{Chlorine Dose (mg/l)} - \text{Chlorine Residual (mg/l)}$$

OR

$$\text{Chlorine Dose (mg/L)} = \text{Chlorine Demand (mg/L)} + \text{Chlorine Residual}$$

Example #1: The chlorine demand of water is 2.1 mg/L. If a chlorine residual of 0.6 mg/L is desired, what is the required chlorine dosage in mg/L?

$$\text{Chlorine Dose (mg/L)} = \text{Chlorine Demand (mg/L)} + \text{Chlorine Residual (mg/L)}$$

$$\begin{aligned} ? \text{ Dose} &= 2.1 + 0.6 \\ &= 2.7 \text{ mg/L} \end{aligned}$$

Example #2: The chlorine dosage is 2.9 mg/L. If the chlorine residual is 0.6 mg/L, what is the chlorine demand in mg/L?

$$\text{Chlorine Demand (mg/L)} = \text{Chlorine Dose (mg/L)} - \text{Chlorine Residual (mg/L)}$$

$$\begin{aligned} ? \text{ Chlorine Demand} &= 2.9 - 0.6 \\ &= 2.3 \text{ mg/L} \end{aligned}$$

Calculating dose and feed rate for a work shift (less than 24 hours)

Problem: You must maintain 0.5 mg/L chlorine residual in the finished water with a chlorine demand of 1.5 mg/L. The pumping rate is 500 gpm. How many pounds of 68% calcium hypochlorite will be fed during the first shift (8 hours)?

Step 1: Calculate the dose using the formula

Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)

$$\begin{aligned} \text{Dose} &= 1.5 + 0.5 \\ &= 2.0 \text{ mg/L} \end{aligned}$$

Step 2: Convert gpm to MGD

$$? \text{MG} = \frac{1 \text{ MG}}{\text{day} \quad 1,000,000 \text{ gal}} \times 500 \frac{\text{gal}}{\text{min}} \times 1440 \frac{\text{min}}{\text{day}} = 0.72 \text{ MGD}$$

Step 3: Use Feed Rate calculation to solve for lbs/day

$$? \text{ lbs/day} = \text{flow} \times \text{dose} \times 8.34 = (0.72)(2.0)(8.34) = 12 \text{ pounds of chlorine is required.}$$

Step 4: Calculate # of pounds of 68% solution needed to achieve Step 3 feed rate.

a) Convert % purity of solution into a decimal:

$$\frac{68\%}{100\%} = 0.68$$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$$\frac{12 \text{ pounds}}{0.68} = 17.6 \text{ pounds of 68\% calcium hypochlorite.}$$

Step 5: Calculate pounds needed for an 8 hour shift

$$? \text{ lbs} = 17.6 \frac{\text{lbs}}{\text{day}} \times \frac{1 \text{ day}}{24 \text{ hrs}} \times \frac{8 \text{ hrs}}{1} = 5.8 \text{ lbs in an 8 hour shift}$$

Practice Problem: You must maintain 0.5 mg/L chlorine residual in the finished water with a chlorine demand of 1.5 mg/L. The pumping rate is 300 gpm. How many pounds of 65% calcium hypochlorite will be fed during **12 hours**?

Step 1: Calculate the dose using the formula

Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)

Dose = _____ mg/L

Step 2: Convert gpm to MGD

?MG / day = _____ x _____ x _____ = _____ MGD

Step 3: Use Feed Rate calculation to solve for lbs/day

? lbs/day = flow x dose x 8.34 = (____)(____)(8.34) = _____ pounds of chlorine is required.

Step 4: Calculate # of pounds of 65% solution needed to achieve Step 3 feed rate.

a) Convert % purity of solution into a decimal:

$\frac{65\%}{100\%} = \underline{\hspace{2cm}}$

b) Then divide the pounds needed (feed rate of 100% pure chemical) by the % purity of the solution (as a decimal).

$\frac{\underline{\hspace{2cm}} \text{ pounds}}{0.65} = \underline{\hspace{2cm}} \text{ pounds of 65\% calcium hypochlorite.}$

Step 5: Calculate pounds needed for 12 hours

?lbs = _____ $\frac{\text{lbs}}{\text{day}}$ x $\frac{1 \text{ day}}{24 \text{ hrs}}$ x $\frac{12 \text{ hrs}}{1}$ = _____ lbs in 12 hours

Calculating Feed Rate and Converting to gallons per day (again)

Problem: How many gallons of 12.5% sodium hypochlorite are required to treat 116,000 gpd with a desired residual of 0.5 mg/L and a chlorine demand of 2.0 mg/L? NOTE: 12.5% sodium hypochlorite = 1.2 lb/gallon available chlorine (“active ingredient” weight).

Remember, the problem asks you to solve for gallons and the available chlorine is given. This is what we called Path 2 in our Tip on Page 3-26.

Step 1: Calculate the dose using the formula

Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)

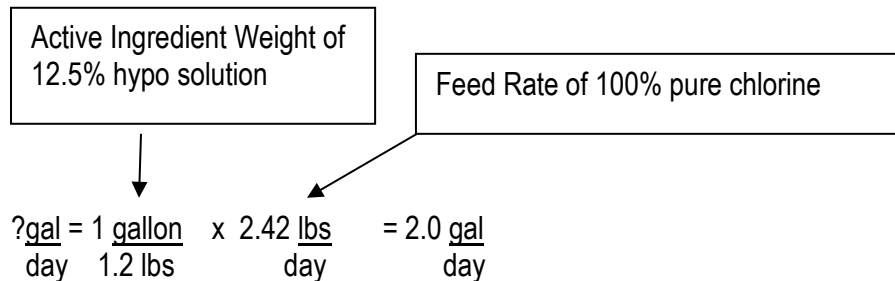
$$\begin{aligned} \text{Dose} &= 2.0 + 0.5 \\ &= 2.5 \text{ mg/L} \end{aligned}$$

Step 2: Convert gpd to MGD

$$\frac{116,000}{1,000,000} = 0.116 \text{ MGD}$$

Step 3: Use Feed Rate calculation to solve for lbs/day

$$? \text{ lbs/day} = \text{flow} \times \text{dose} \times 8.34 = (0.116)(2.5)(8.34) = 2.42 \text{ pounds of chlorine is required.}$$

Step 4: Use “active ingredient” weight with unit cancellation steps to convert lbs/day to gallons/day

NOTE: When you are given the “active ingredient” weight of a solution to solve a feed rate problem, you do not need to use the % purity factor because it was used in the weight calculation.

CALCULATIONS: CHLORINE DEMAND OR DOSE

Practice Problem: How many gallons of 12% sodium hypochlorite are required to treat 150,000 gpd with a desired residual of 0.8 mg/L and a chlorine demand of 0.6 mg/L? NOTE: 12% sodium hypochlorite – 1.2 lb/gallon available chlorine (“active ingredient” weight).

Step 1: Calculate the dose using the formula

Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)

Dose = _____ mg/L

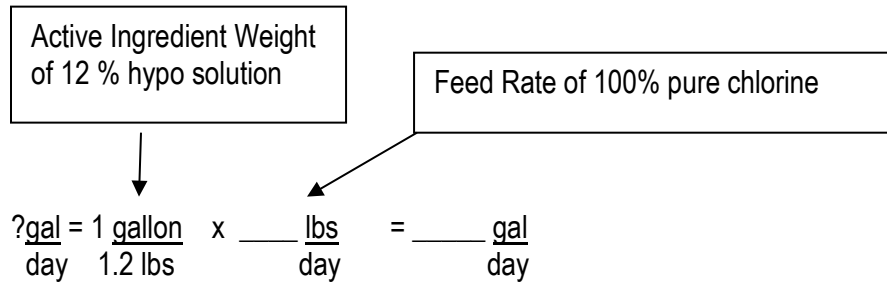
Step 2: Convert gpd to MGD

$\frac{150,000}{1,000,000} = \text{_____ MGD}$

Step 3: Use Feed Rate calculation to solve for lbs/day

? lbs/day = flow x dose x 8.34 = (____)(____)(8.34) = _____ pounds of chlorine is required.

Step 4: Use “active ingredient” weight with unit cancellation steps to convert lbs/day to gallons/day



CT



CT – The product of residual disinfectant concentration (C) measured in mg/L in a representative sample of water prior to the first customer, and disinfection contact time (T) in minutes; that is “C” x “T.”

$$\text{CT} = \text{disinfectant concentration} \times \text{contact time} = \text{C (mg/L)} \times \text{T (minutes)}$$

Therefore, the *units of CT are expressed in mg-min/L.*

Example #1: If a free chlorine residual of 1.8 mg/L is measured at the entry point of the system, after 120 minutes of detention time in the clearwell, what is the CT value in mg-min/L?

$$\begin{aligned} \text{CT} &= \text{disinfectant concentration} \times \text{contact time} \\ 1.8 \text{ mg/L} &\quad \times \quad 120 \text{ minutes} = 216 \text{ mg-min/L} \end{aligned}$$

Example #2: If a free chlorine residual of 3.0 mg/L is measured at the end of the clearwell after 5 hours of detention time, what is the CT value in mg-min/L?

Step 1: Convert 5 hours of detention time to minutes (CT must be in minutes)

$$\text{Min} = \frac{60 \text{ min}}{\text{hr}} \times 5 \text{ hr} = 300 \text{ minutes}$$

Step 2: Insert disinfectant residual concentration and contact time (in minutes) into CT equation and multiply the values.

$$\begin{aligned} \text{CT} &= \text{disinfectant concentration} \times \text{contact time} \\ \text{CT} &= 3.0 \text{ mg/L} \times 300 \text{ min} = 900 \text{ mg-min/L} \end{aligned}$$

Practice Problem: If a free chlorine residual of 2.5 mg/L is measured at the end of the clearwell after 4 hours of detention time, what is the CT value in mg-min/L?

Step 1: Convert detention time from hours to minutes.

? Min =

Step 2: Insert disinfectant residual concentration and contact time (in minutes) into CT equation and multiply the values.

CT = _____ mg/L x _____ minutes = _____ mg-min/L

Calculating Theoretical Detention Time

$$\text{Theoretical Detention Time (minutes)} = \frac{\text{Volume of Tank (gallons)}}{\text{Influent Flow (gpm)}}$$

Notice the units:

Volume = gallons
Time = minutes

All calculations must be converted to those units before you do the math when using it in the CT calculation.

Example #3: A sedimentation tank holds 50,000 gallons and the flow into the plant is 500 gpm. What is the detention time in minutes?

$$\text{Detention Time (mins)} = \frac{\text{Volume}}{\text{Flow}} = \frac{50,000 \text{ gallons}}{500 \text{ gpm}} = 100 \text{ minutes}$$

Practice Problem: What is the detention time (in min) of a tank that has a volume of 150,000 gallons with a plant flow rate of 2.5 MGD? (hint convert the flow rate to gpm, then plug into DT)

Step 1: Convert flow rate from MGD to gpm

$$\frac{? \text{ gal}}{\text{min}} = \frac{1,000,000 \text{ gal}}{1 \text{ MG}} \times \frac{2.5 \text{ MG}}{\text{day}} \times \frac{1 \text{ day}}{1440 \text{ mins}} = \frac{\text{gal}}{\text{mins}}$$

Step 2: Calculate detention time using the formula:

$$\text{Detention Time (mins)} = \frac{\text{Volume}}{\text{Flow}} = \frac{\text{gallons}}{\text{gpm}} = \text{minutes}$$

Now we are ready to combine this information with the CT calculation in a problem.

Example #4: A plant is set at a flow rate of 5 MGD. Water enters into a clearwell that has a volume of 75,000 gallons. The chlorine residual of the outlet end of the tank is 1.2 mg/L. What is the CT in mg-min/L?

Step 1: Convert flow rate from MGD to gpm

$$\frac{? \text{ gal}}{\text{min}} = \frac{1,000,000 \text{ gal}}{1 \text{ MG}} \times \frac{5 \text{ MG}}{\text{day}} \times \frac{1 \text{ day}}{1440 \text{ mins}} = \frac{3472 \text{ gal}}{\text{mins}}$$

Step 2: Calculate detention time using the formula:

$$\text{Detention Time (mins)} = \frac{\text{Volume}}{\text{Flow}} = \frac{75,000 \text{ gallons}}{3472 \text{ gpm}} = 21 \text{ minutes}$$

Step 3: Insert disinfectant residual concentration and contact time (in minutes) into CT equation and multiply the values.

$$\text{CT} = 1.2 \text{ mg/L} \times 21 \text{ minutes} = 25 \text{ mg-min/L}$$

Practice Problem: A plant is set at a flow rate of 3 MGD. Water enters into a clearwell that has a volume of 50,000 gallons. The chlorine residual of the outlet end of the tank is 1.6 mg/L. What is the CT in mg-min/L?

Step 1: Convert flow rate from MGD to gpm

?gal =
min

Step 2: Calculate detention time using the formula:

$$\text{Detention Time (mins)} = \frac{\text{Volume}}{\text{Flow}} \quad \frac{\text{gallons}}{\text{gpm}} = \text{minutes}$$

Step 3: Insert disinfection residual concentration and contact time (in minutes) into CT equation and multiply the values.

$$\text{CT} = \text{mg/L} \times \text{minutes} = \text{mg-min/L}$$



Key points for Chemical Feed Dosage, Chlorine Demand and CT Calculations

- Remember to perform math calculations using the order of operation steps listed in this unit.
- You can use unit cancellation steps to solve for the units you are seeking. Be sure to begin with the numerator unit and cancel unwanted units until only the unknown denominator units remain.
- The Davidson Pie diagram can be used to solve for feed rate (lbs or lbs/day), flow (MGD) or dosage (mg/L) by using the following formulas:
 1. Feed Rate, lbs/day = Flow (MGD) or Volume (MG) x Dosage (mg/L) x 8.34 (which is the density of water)
 2. Flow (MGD) = $\frac{\text{Feed Rate (lbs/day)}}{[\text{Dosage (mg/L)} \times 8.34]}$
 3. Dosage (mg/L) = $\frac{\text{Feed Rate (lbs/day)}}{[\text{Flow(MGD)} \times 8.34]}$
- In order to use any of these formulas, all flows or volumes must be converted to either million gallons per day (MGD) or million gallons (MG).
- If you are calculating a feed rate for a solution that is not 100% pure chlorine (like sodium hypochlorite and calcium hypochlorite), remember to:
 - Calculate the feed rate as if it's 100% pure,
 - Then divide the feed rate of 100% pure chemical by the percent purity of the solution (as a decimal).
- Specific gravity is used to calculate the “active ingredient” weight of a solution which provides you with the available lbs of chlorine/gallon of solution. Once you have the available lbs of chlorine/gallon of solution (i.e., “active ingredient” weight), you can use unit cancellation to determine the number of gallons you need to feed.
- You can use the following equation to calculate chlorine dose, chlorine demand or chlorine residual.
 - Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)
- To calculate CT values, use the following formulas:
 - CT = disinfectant concentration x contact time = C (mg/L) x T (minutes)
 - Theoretical Detention Time (minutes) = $\frac{\text{Volume of Tank (gallons)}}{\text{Influent Flow (gpm)}}$
 - Remember to convert flow into **gpm** and volume into **gallons**
 - Remember to convert hours into **minutes** before using the formulas.

**Exercise for Unit 3 – Chemical Feed Dosage, Chlorine Demand and CT Calculations**

1. In order to use the Feed Rate formula which is $\text{lbs/day} = \text{Flow or Volume} \times \text{Dosage} \times 8.34$, name the units of measurement for the flow or volume:
 - a) MGD or MG
 - b) gpm or gallons
 - c) gpd or gallons
 - d) All of the above units can be used

2. If you have calculated the feed rate for a solution as if it's 100% pure; but, your solution is a 65% calcium hypochlorite, what value do you use to represent the percent purity (as a decimal)? In other words, what value are you dividing by?
 - a) 65
 - b) 6.5
 - c) 0.65
 - d) 0.0065

3. You have determined that you need to feed 100 lbs/day of chlorine. You are using 15% sodium hypochlorite which provides 1.2 lbs/gal available chlorine. In order to convert the "lbs/day" feed rate into "gallons/day," what math step do you use?
 - a) $100 \text{ lbs/day} \times 1.2 \text{ lbs/gal}$
 - b) $100 \text{ lbs/day} \times 0.15$
 - c) $100 \text{ lbs/day} \div 1.2 \text{ lbs/gal}$
 - d) $100 \text{ lbs/day} \div 0.15$

4. When calculating a CT value, what units are used in the detention time calculation?
 - a) $\text{Volume (MG)} \div \text{Flow (gpm)}$
 - b) $\text{Volume (Gal)} \div \text{Flow (gpm)}$
 - c) $\text{Volume (MG)} \div \text{Flow (MGD)}$
 - d) $\text{Volume (Gal)} \div \text{Flow (MGD)}$

Additional practice math problems are located in the Appendix.

Reference

Joanne Kirkpatrick Price, *Basic Math Concepts for Water and Wastewater Plant Operators*.

Unit 4 – Chemical Feed

Learning Objectives

- Explain the disinfection regulatory requirements.
- Explain breakpoint chlorination.
- Identify chemical feed equipment and explain important operation and maintenance considerations

Regulatory Requirements

All community water systems (CWSs) are required to provide continuous disinfection. Also, CWSs must meet the disinfection byproducts MCLs. The following regulatory requirements address both disinfection and their associated byproducts with treatment technique requirements, MCLs, and MRDLs.

Surface Water Supplies

- The continuous filtration and disinfection process must achieve 99.9 percent (3 log) inactivation of *Giardia* cysts and 99.99 percent (4 log) inactivation of enteric viruses. Chlorination equipment must be capable of maintaining a chlorine residual, which achieves a minimum of 1.0 log *Giardia* cyst inactivation following filtration.
- This must be determined by CT factors and measurement methods established by EPA. Refer to EPA's *Guidance Manual for Compliance with the Filtration and Disinfection Requirements for Public Water Systems Using Surface Water Sources*.

Groundwater Supplies

- Community groundwater systems are required to provide continuous disinfection and at least 4-log treatment of viruses (99.99% removal and/or inactivation).
- Community groundwater systems are required to maintain at each groundwater entry point a minimum residual disinfection concentration approved by DEP to provide 4-log treatment of viruses.
- This applies to groundwater supplies **not** under the influence of surface water intrusion.
- A minimum of 20 minutes of contact time must be provided.

Entry Point Minimum Chlorine Residual Requirements:

- **For all surface water systems:**
 - Surface water systems using filtered surface water (SW) or groundwater under the direct influence of SW (GUDI) sources must maintain a minimum disinfectant residual at the entry point (EP) of at least 0.20 mg/L. The previous minimum residual was 0.2 mg/L. The EP residual may not be less than 0.20 mg/L for more than 4 hours.
- **For groundwater systems (all Community Water Systems and Noncommunity systems permitted for 4-log disinfection):**
 - Minimum free chlorine entering the distribution system (i.e., **entry point**) no less than 0.40 mg/L or its equivalent as approved by DEP or other minimum residual approved by DEP to provide 4-log treatment of viruses. The EP residual may not be less than 0.40 mg/L for more than 4 hours.

Distribution System Minimum Chlorine Residual Requirements

- The following systems must maintain a minimum disinfectant residual concentration in the distribution system of at least 0.2 mg/L (unless a higher number has been specified in a permit):
 - All community water systems
 - All non-transient noncommunity water systems using a chlorine disinfectant
 - Transient noncommunity water systems with filtration of surface water or 4-log disinfection of a groundwater source.
- Disinfectant residual measurements must be collected at representative locations at the same time and location(s) as coliform samples.
- Disinfectant residual measurements must also be conducted at least once/week. Water systems that do not collect at least 1 coliform sample each week will need to conduct additional disinfectant residual measurements each week that a coliform sample is not collected.
- Any location that has a disinfectant residual less than 0.2 mg/L in a month must also be sampled the following month.

Maximum Contaminant Levels

Chlorine added to water containing organic and inorganic chemicals reacts with these materials to form chlorine compounds. Maximum Contaminant Levels (MCLs) have been established for these disinfection byproduct compounds.



Maximum contaminant levels (MCLs) regulations are issued by the USEPA and in Pennsylvania by the DEP. The MCLs list a variety of organic and inorganic chemicals, disinfection by products (DBPs), radionuclides, microbiological contaminants, and turbidity levels that pertain to drinking water. Some example MCLs are listed below.

Maximum Residual Disinfectant Level



The maximum residual disinfectant level (MRDL) is the maximum permissible level of a disinfectant added for water treatment that may not be exceeded at the consumer's tap without an unacceptable possibility of adverse health effects.

<u>Contaminant</u>	<u>MCL or MRDL</u>
Total Trihalomethanes (TTHMs), a DBP	0.080 mg/L
Haloacetic Acids (HAA5), a DBP	0.060 mg/ L
Bromate, a DBP	0.010 mg/ L
Chlorite, a DBP	1.0 mg/L
Benzene	0.005 mg/ L
Diquat	0.02 mg/ L
Chlorine (as Cl ₂)	4.0 mg/ L as maximum residual disinfectant level (MRDL)

Secondary contaminant levels are also included since these may affect taste and odor. Some examples are shown below.

<u>Contaminant</u>	<u>MCL</u>
• Sulfate	250 mg/ L
• Iron	0.3 mg/ L
• pH	6.5 – 8.5

Minimizing Total Trihalomethane (THM) Formation

Drinking water systems can reduce THM formation in several ways:

- Reduce the organic material before chlorinating the water. Treatment techniques, such as coagulation, sedimentation, and filtration can remove most of the organic materials. However, activated carbon can be used to remove greater amounts of organic material than can be removed by other techniques.
- Optimize chlorine usage.
- Change the point of chlorine addition in the treatment series. If the point of chlorine addition is moved to a location **after** sedimentation or filtration, THM production can be reduced as these processes remove part of the organic material.
- Use alternative disinfection methods. Using a mixture of chlorine and ammonia (chloramine) reduces THM formation.

One-Hour Reporting Violations or Situations

A public water supplier shall report the circumstances to the Department within an hour of discovery for the following:

- A primary MCL or an MRDL has been exceeded or treatment technique requirement has been violated.
- A sample result that requires check samples to be collected.
- Circumstances exist which may adversely affect the **quality** or **quantity** of drinking water.
- Any sample result that is *E. Coli* positive.

Chlorination Mechanics and Terminology

- The exact mechanism of chlorine disinfection action is not fully known.
 - One theory is that chlorine directly destroys the bacterial cell.
 - Another theory is that chlorine inactivates the enzymes which enable the cells to use food, thus starving the organisms.



Chlorine demand is the amount of chlorine required to react with all the organic and inorganic material.

- In practice, the chlorine demand is the difference between the amount of chlorine added and the amount remaining after a given contact time.
- Some reactive compounds have disinfecting properties while others do not.



Chlorine residual (or Total Chlorine) is the total of all compounds with disinfecting properties and any remaining free chlorine.

Total Chlorine Residual (mg/L) = Combined Chlorine Forms (mg/L) + Free Chlorine (mg/L)

Combined Chlorine: When chlorine is added to water, some of the chlorine reacts with nitrate to form chlorine-ammonia compounds. This is referred to as combined chlorine.

Free Chlorine: Chlorine available to inactivate disease-causing organisms

- The presence of measurable chlorine residual indicates that all chemical reactions have been satisfied and that sufficient chlorine is present to kill or inactivate the microorganisms.
- **Note:** The chlorine residual is not impacted by changes in water temperature.



Chlorine dose is the amount of chlorine needed to satisfy the chlorine demand plus the amount of chlorine residual needed for disinfection:

Chlorine Dose (mg/L) = Chlorine Demand (mg/L) + Chlorine Residual (mg/L)

Breakpoint Chlorination



Breakpoint chlorination is the addition of chlorine until all chlorine demand has been satisfied. At this point, further additions of chlorine will result in a free chlorine residual that is directly proportional to the amount of chlorine added beyond the breakpoint. Breakpoint chlorination determines how much chlorine is required for disinfection.

- A residual in the form of free available residual chlorine needs to be provided. This has the highest disinfecting ability.

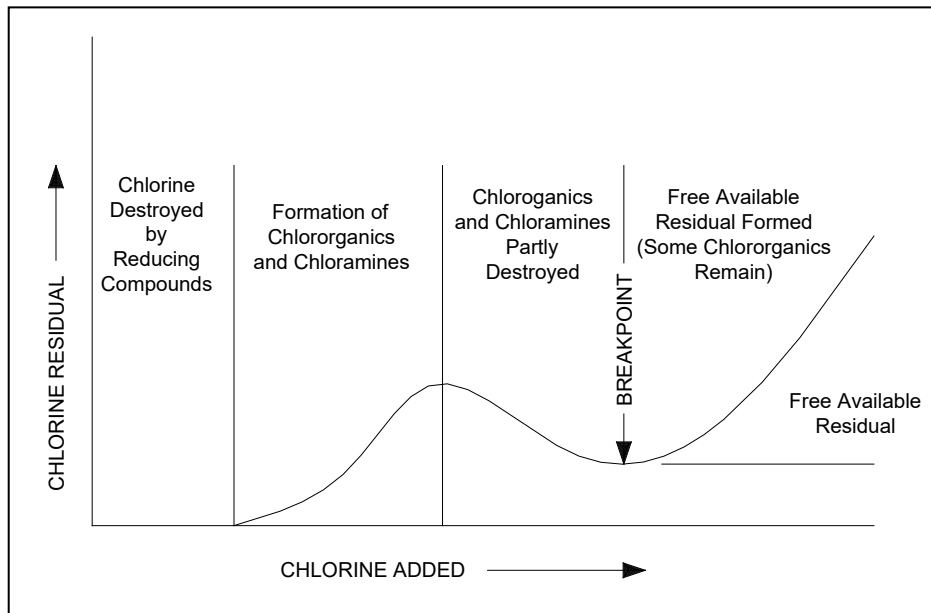
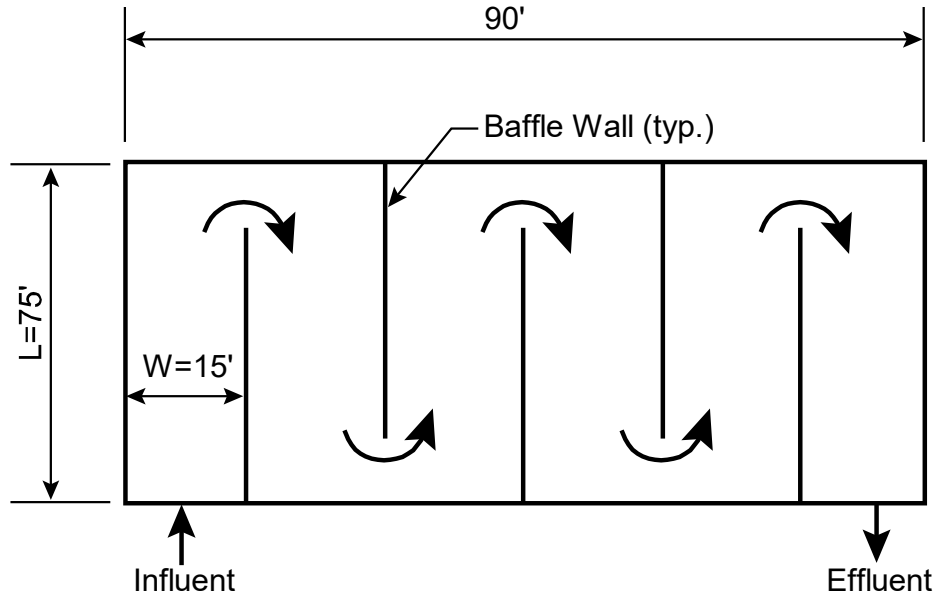


Figure 4.1 Breakpoint Chlorination Curve

Chlorine Contact Tank Design

The chlorine contact basin design is of critical importance to maximize the effective detention time through the basin, and minimize short-circuiting. Baffling within the basins help minimize short-circuiting and aids in mixing. Note that the length to width ratio of 25:1 or greater should result in effective storage or baffling efficiencies of 70% or greater.



$$L:W = \geq 25$$

for 70% Baffling Efficiency

$$\frac{6 \times 75}{15} = 30 \geq 25$$

Figure 4.2 Chlorine Contact Basin.

Sodium Hypochlorite

- Sodium hypochlorite is supplied only in solution form.
- It is ready for use as received.
- Sodium hypochlorite is usually fed neat (i.e., at the strength received) if 12.5% strength or less. If the solution strength is greater than 12.5%, it may need to be diluted.

Figure 4.3 shows a typical bulk sodium hypochlorite feed system schematic diagram.

Figure 4.4 depicts a schematic diagram of a typical sodium hypochlorite drum feed system.

HYPOCHLORITE FEED EQUIPMENT, OPERATION AND MAINTENANCE

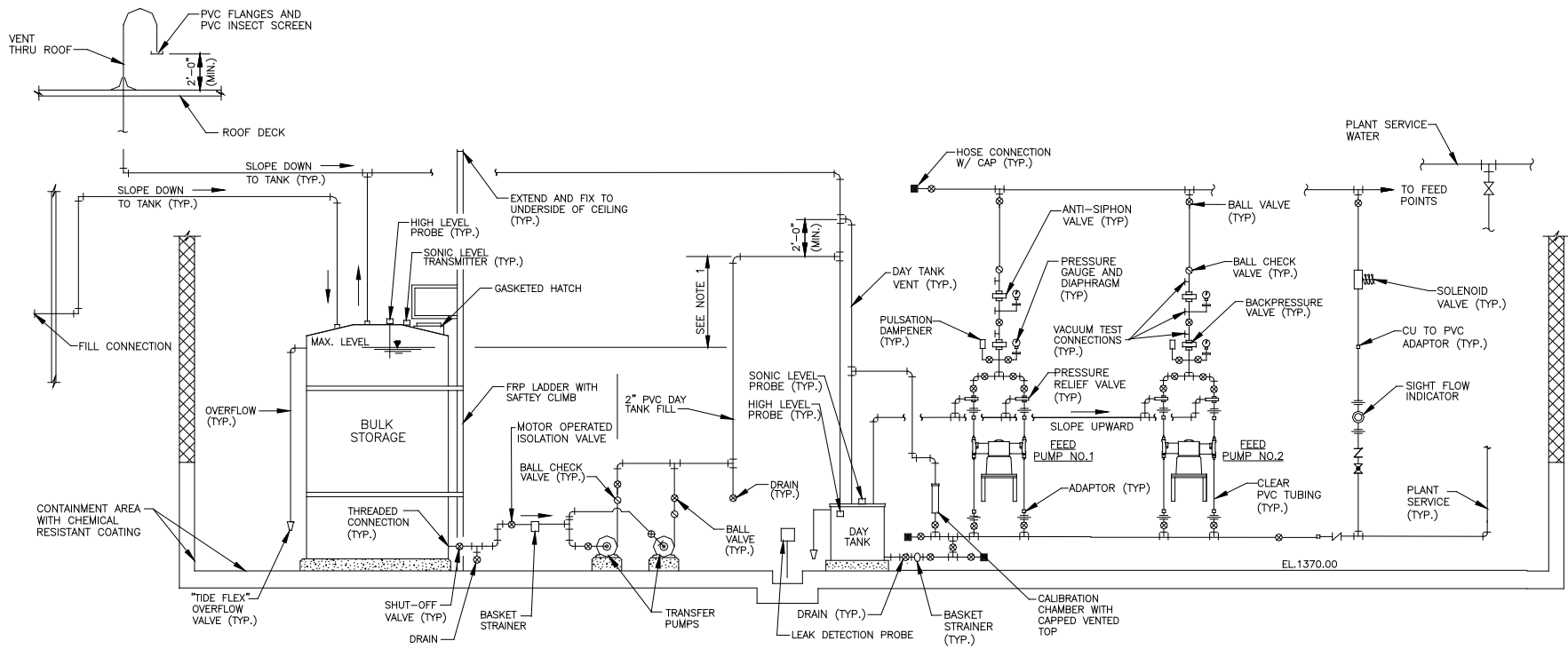


Figure 4.3 Typical Bulk Sodium Hypochlorite Feed System Schematic

HYPOCHLORITE FEED EQUIPMENT, OPERATION AND MAINTENANCE

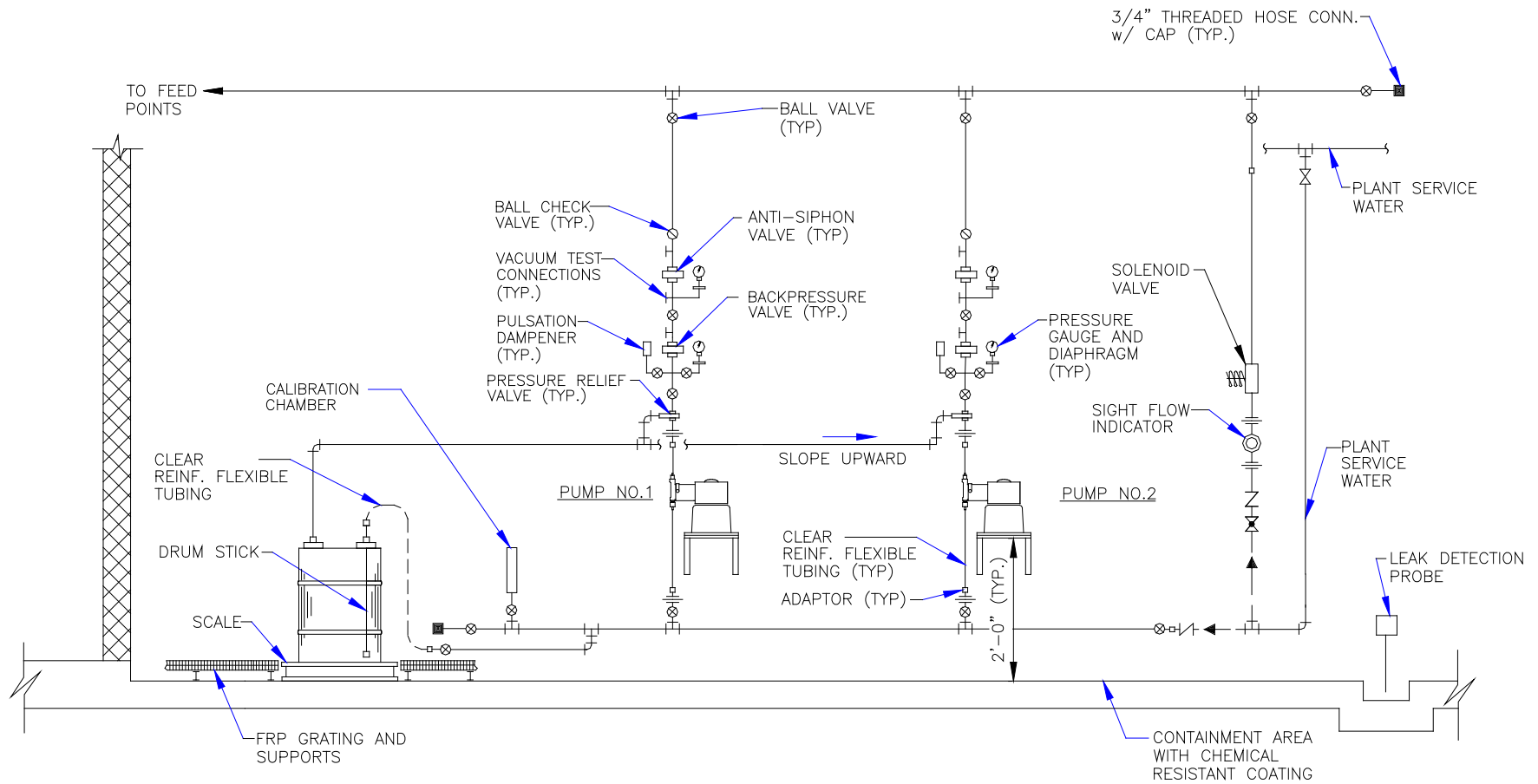


Figure 4.4 Typical Sodium Hypochlorite Drum Feed System Schematic

Hypochlorite Feed Equipment

With the exception of the type of chemical storage/solution preparation, the equipment for feeding the two forms of hypochlorite is similar. The major components and the purpose of each follow.

Storage/Solution Preparation

- The bulk storage tank provides a container for the storage of a 30-day minimum supply of sodium hypochlorite solution. The bulk storage tank should provide a minimum volume of 110% of the maximum chemical delivery quantity to minimize the potential for overflowing the tank during chemical deliveries.
- The solution preparation tank provides containers for preparation of a minimum 1-day supply of calcium hypochlorite solution. Two containers are suggested: one in service and one with solution aging to facilitate settling of insoluble compounds.

Scales

- Scales provide an indication of the quantity of chemical remaining, which is useful for chemical inventory control.

Transfer Pumps

- Transfer pumps are used for the transfer of sodium hypochlorite from the bulk storage tank to day tanks.

Day Tank

- The day tank stores daily chemical required for delivery by feeders, and also monitors chemical usage and provides inventory control.

Chemical Feeder

- The chemical feeder uses positive displacement diaphragm pumps such as the mechanical diaphragm type, the hydraulic diaphragm type, or an electronic solenoid pump. Some chemical feeders are vacuum eductor-type feeders.
- The calibration cylinder or calibration chamber consists of a graduated cylinder typically located on the suction side of the pump. It is used for accurate determination of the pump's feed rate.

Chemical Feed Piping

- The backpressure valve maintains a constant backpressure on the feed pump discharge.
- The anti-siphon valve prevents backsiphonage of process water into the chemical feed system.
- The pressure relief valve limits the discharge pressure of the feed pump and protects the feed piping.
- The isolation valves permit maintenance of various system components without the need to remove the entire feed system from operation.

Calcium Hypochlorite

Most plants manually prepare stock batch solutions. Figure 4.5 depicts a typical Calcium Hypochlorite Feed System.

- Usually a one-day supply of 1 to 3 % strength is prepared.
- The proper quantity of HTH is dissolved in a sufficient quantity of water.
- The batch solution should be prepared in a separate container.
 - Allow the solution to stand for 8 to 24 hours before actual use.
 - Pour or siphon the clear solution into the day tank supplying the feeder.

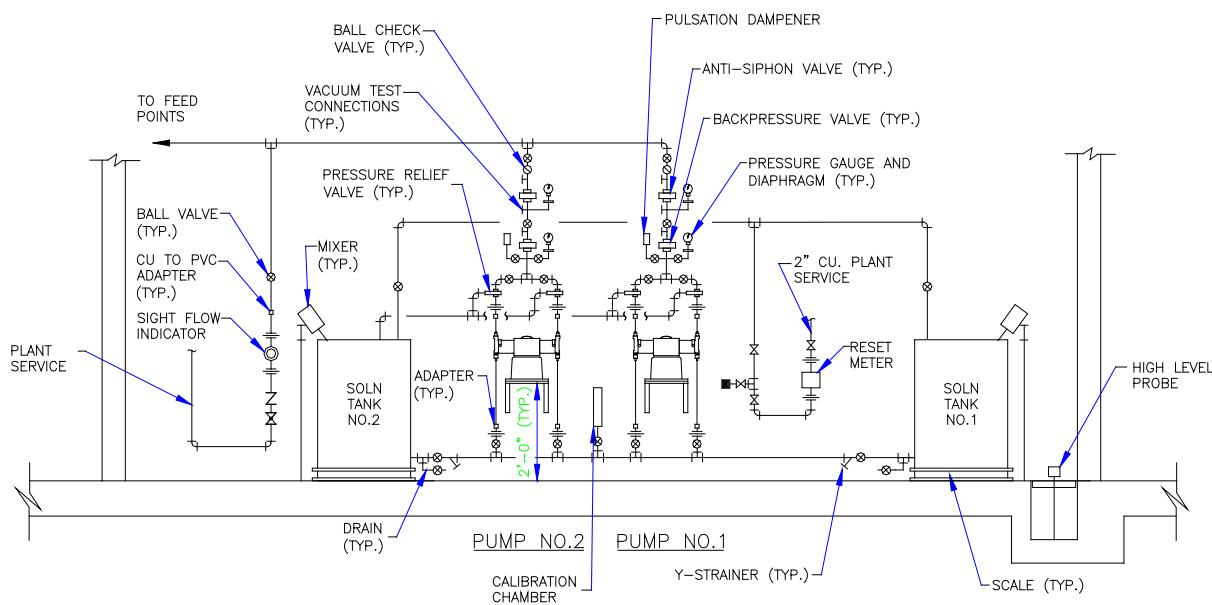


Figure 4.5 Typical Calcium Hypochlorite Feed System

On-Site Generated Sodium Hypochlorite

Typical sodium hypochlorite generation process layout is shown in the following figure.

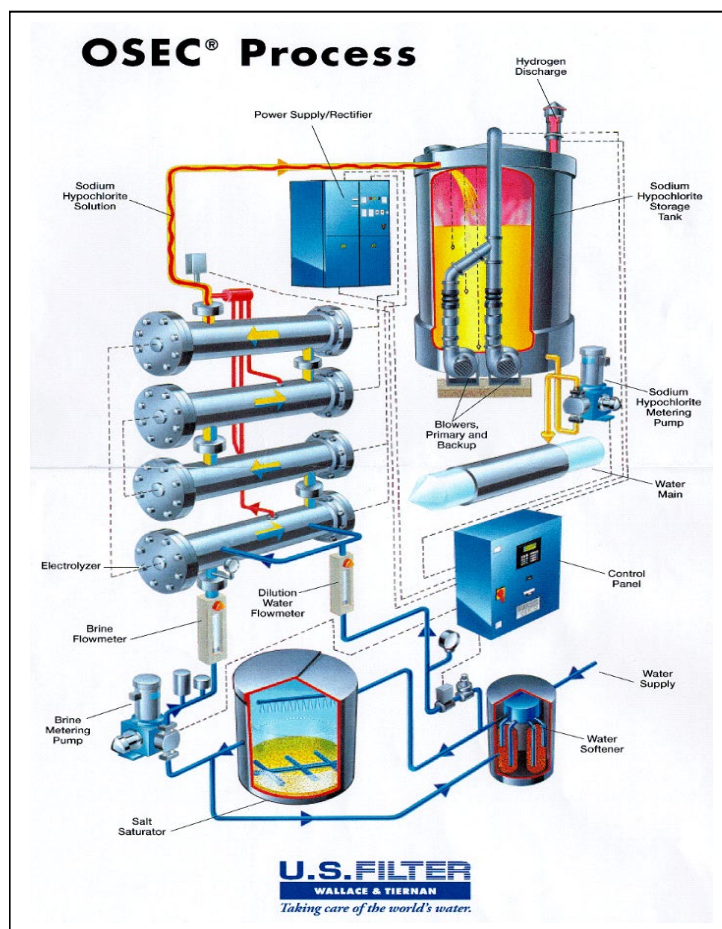


Figure 4.6 On-Site Hypochlorite Generation Process¹

System Components

- The brine storage vessel stores salt and softened water where concentrated brine solution is made and stored. (Noted as Salt Saturator in Figure 4.6.)
- The water softener reduces the hardness of the water supply to less than 17 mg/l total hardness.
- The hypochlorite generators are reaction vessels where electrical charge is applied to the brine solution, resulting in the formation of sodium hypochlorite. (Noted as Electrolyzer in Figure 4.6)
- The control system or panel controls the hypochlorite production cycle.
- The storage tanks store hypochlorite solution. Hydrogen is released from the process.

Chemical Feed Settings

Feed Rate is the quantity or weight of chemical delivered from a feeder over a given period of time. The chemical feed pump must be calibrated to deliver the selected dosage. A feed rate can have different units of expression, such as lb/hr, lbs/day, mL/min, or gal/day. Often, determining a feed rate involves time and weight conversions.

- Feed pumps are calibrated with the use of a pump calibration curve.
- A new pump calibration curve should be constructed:
 - At least once per year.
 - If troubleshooting points to the need for a new pump calibration.
 - If any maintenance is performed on the pump.

Steps in Developing a Pump Calibration Curve

Step 1: Determine actual feed pump output.

- Operate feed pump over full operating range
- Determine actual pump output

LIQUID FEED PUMP CALIBRATION TABLE

% Stroke: _____

PUMP SETTING	VOLUME (mL)	TIME (min)	FEED RATE (mL/min)
20			
40			
60			
80			
100			

Here's an example of a completed liquid feed pump calibration table.

Pump Setting (%)	Alum Pumped (ml)	Time (sec)	Feed Rate (ml/min)	Feed Rate (gal/min)
0	0.0	30	0.00	0.000
20	65.6	55	71.56	0.019
40	141.9	59	144.31	0.038
60	249.1	61	245.02	0.065
80	195.2	32	366.00	0.097
100	267.4	35	458.40	0.121

Figure 4.7 Liquid Feeder Pump Calibration Table

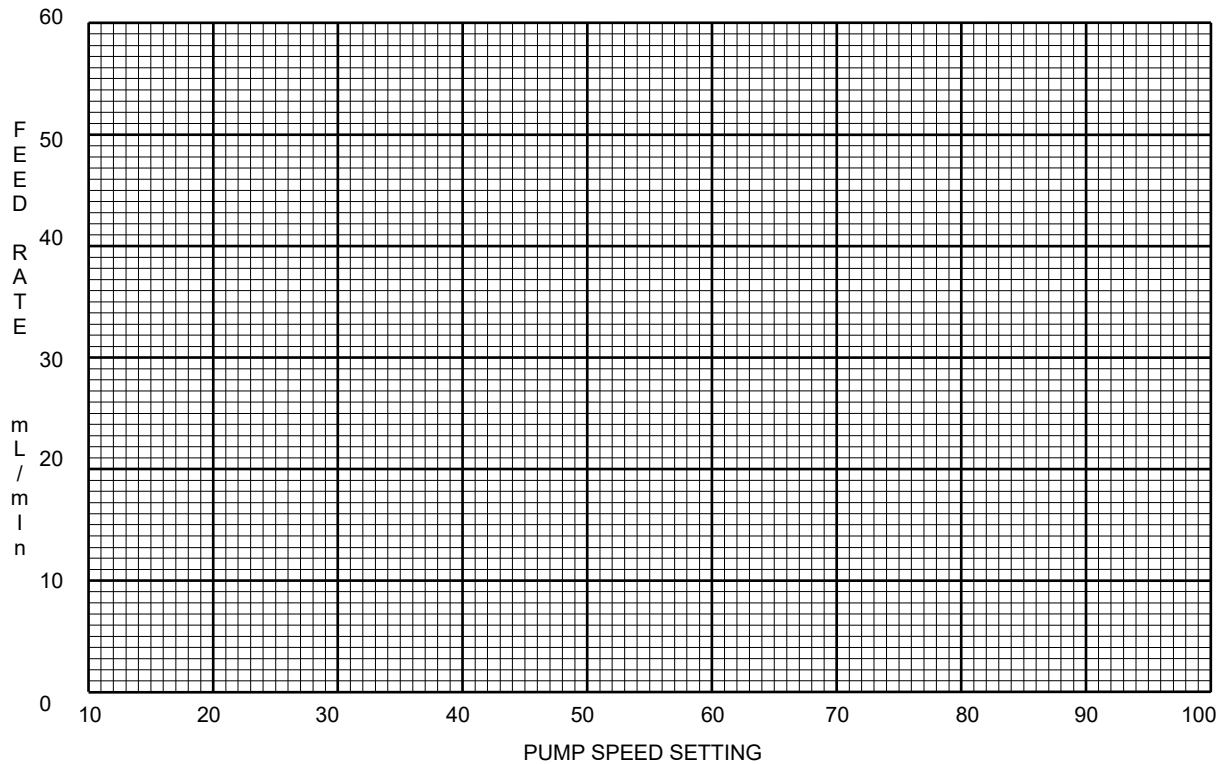
HYPOCHLORITE FEED EQUIPMENT, OPERATION AND MAINTENANCE

Step 2 – Develop feed pump calibration curve.

- Plot each **Feed Rate (mL/min)** vs. **Pump Speed setting** on the graph.
- Connect each of the points together with a straight line.

PUMP: _____ PUMP CALIBRATION CURVE DATE: _____

% Stroke: _____



Here's an example of a Dry Feeder Calibration Curve.

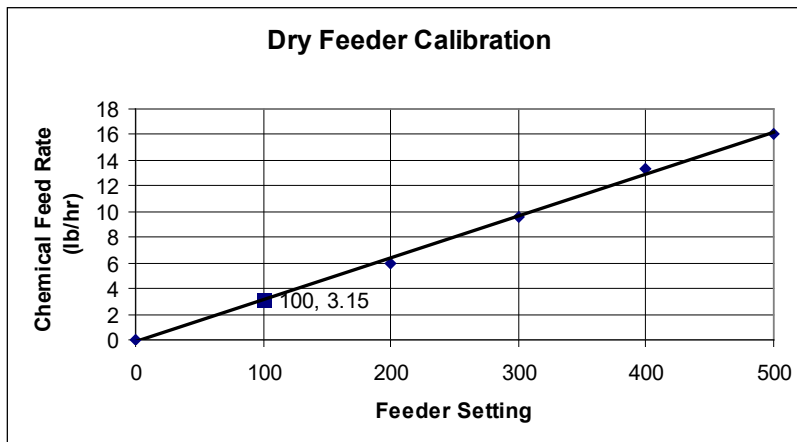


Figure 4.8 - Feeder Calibration Curve

Hypochlorite Feed System Operation

Normal Operation

Normal operation of the hypochlorite feed system requires regular observation of the facilities and equipment, and a regular preventative maintenance program recommended in the manufacturer's specifications. Exact operating procedures will depend on the type of hypochlorite in use at the facility and on the equipment that is installed at the facility.

The general procedure is as follows:

Daily

- Visually inspect the storage and feed areas.
- Verify operation of chemical transfer pumps.
- Read scales, charts or meters at same time each day to determine actual chemical usage. Be sure to record readings for water pumped and actual chemical usage.
- Check chlorine residual in the system and adjust chlorine feed rate as necessary. Verify the proper operation of chemical feed pumps.
- Prepare calcium hypochlorite solution as necessary.
- Look for leaks.

Weekly

- Check the chemical dose by verifying proper pump calibration and computing the actual weekly chemical dosage.
- Check the entire system for problems.
- Clean the area.
- Verify the chemical supply on-hand and order as necessary.

Monthly

- Clean and lubricate equipment in accordance with manufacturer's recommendations.

Abnormal Operation

In the event of an abnormal operation, be sure to inform your supervisor of the problem.

Low Chlorine Residual

- Determine the actual chlorine residual in the laboratory and compare with the residual analyzer reading. Then recalibrate the analyzer appropriately.
- If the analyzer is properly calibrated, check the following:
 - The sample pump operation.
 - The feed pump operation.
 - The control system, if it is on automatic control. Operate in manual control mode if necessary.
 - The water chlorine demand. Increase the feed rate if necessary. Add additional feed pumps on line if demand is higher than can be provided by a single unit.

Chemical Pump Not Operating

- Verify the hypochlorite supply availability.
- Start the spare pump to maintain system chlorination.
- Check the electrical supply equipment, such as electrical connections, circuit breakers and control equipment.
- Check for blockages:
 - In the solution tank.
 - In the valves, both manual and electrically controlled.
 - In feed lines such as the feed pump suction lines and feed pump discharge lines.



Key points for Unit 4 – Chemical Feed.

- Public water suppliers have both entry point and distribution chlorine residual requirements as follows:
 - **Entry Point Minimum Chlorine Residual Requirements:**
 - **For surface water systems:**
 - Surface water systems using filtered surface water (SW) or groundwater under the direct influence of SW (GUDI) sources **must maintain a minimum disinfectant residual at the entry point (EP) of at least 0.20 mg/L.** The previous minimum residual was 0.2 mg/L.
 - **For groundwater systems:**
 - **Minimum free chlorine entering the distribution system (i.e., entry point) no less than 0.40 mg/L or its equivalent as approved by DEP or other minimum residual approved by DEP to provide 4-log treatment of viruses.**
 - **Distribution System Chlorine Residual Requirements:**
 - The following systems must maintain a minimum disinfectant residual concentration in the distribution system of at least 0.2 mg/L (unless a higher number has been specified in a permit):
 - All community water systems
 - All non-transient noncommunity water systems using a chlorine disinfectant
 - Transient noncommunity water systems with filtration of surface water or 4-log disinfection of a groundwater source.
 - Disinfectant residual measurements must be collected at representative locations throughout the distribution system on a weekly basis.
 - **Maximum residual disinfectant level requirements within the distribution system for all system types:** Surface water and groundwater suppliers are required comply with a maximum residual disinfectant level no greater than 4.0 mg/L as chlorine. Compliance is based on a running annual average, computed quarterly, of monthly averages of all samples collected within the distribution system.
- Drinking water systems can reduce THM formation in several ways:
 - Reduce the organic material before chlorinating the water. Treatment techniques, such as coagulation, sedimentation, and filtration can remove most of the organic materials. However, activated carbon can be used to remove greater amounts of organic material than can be removed by other techniques.
 - Optimize chlorine usage.
 - Change the point of chlorine addition in the treatment series. If the point of chlorine addition is moved to a location **after** sedimentation or filtration, THM production can be reduced as these processes remove part of the organic material.

- Use alternative disinfection methods. Using a mixture of chlorine and ammonia (chloramine) reduces THM formation.
- Chlorine demand is the amount of chlorine required to react with all the organic and inorganic material.
- Chlorine residual is the total of all chlorine compounds with disinfecting properties and any remaining free chlorine.
- Chlorine dose is the amount of chlorine needed to satisfy the chlorine demand plus the amount of chlorine residual needed for disinfection.
- Breakpoint chlorination is the addition of chlorine until all chlorine demand has been satisfied. It is used to determine how much chlorine is required for disinfection. At this point, further additions of chlorine will result in a free chlorine residual that is directly proportional to the amount of chlorine added beyond the breakpoint.
- The chlorine contact basin design is of critical importance to maximize the effective detention time through the basin, and minimize short-circuiting. Groundwater systems now required to provide 4-log treatment of viruses may need to install baffles within the chlorine contact tanks to increase detention time and reduce short-circuiting.
- Sodium hypochlorite is supplied in solution form and typically it is used as received.
- Sodium hypochlorite can be generated on site by using special equipment to supply an electric charge to a brine solution.
- After a chlorine dosage has been calculated, it's important to develop a chlorine feed pump calibration curve that determines the chlorine delivery feed rate as it relates to the pump setting. This calibration process assures that the pump is delivering the proper chlorine dosage.
- Hypochlorite feed system equipment should be inspected daily for proper operation. Follow the maintenance recommendations that the pump manufacturer provides.
- Check the chlorine residual in the system daily and adjust chlorine feed rate as necessary.



Exercise for Unit 4 – Chemical Feed

1. Surface water supplies must provide a minimum disinfection residual of _____ mg/L at the entry point.
2. Weekly chlorine residual samples are to be taken at representative points within the distribution system.
 - a) True
 - b) False
3. The _____ is the maximum permissible level of a disinfectant added for water treatment that may not be exceeded at the consumer's tap without an unacceptable possibility of adverse health effects.
4. List one way a water supplier can reduce THM formation:

5. Explain what breakpoint chlorination is.

6. The _____ chlorination curve can be used to determine how much chlorine is required for disinfection.
7. Chlorine dose = _____ (mg/L) + _____ (mg/L).
8. A _____ tank stores daily amounts of chemical required for delivery by feeders.
9. Calcium hypochlorite solutions are typically prepared with a _____ to _____ % strength.
10. A pump calibration curve plots feed rate delivery versus the pump _____.
11. In the event of an abnormal feed equipment operation, be sure to inform your _____ about the problem.

¹ Courtesy of U.S. Filter, Wallace & Tiernan Products (18 June 2003).

Appendix: Safety Data Sheet (SDS) Example

Sodium Hypochlorite, 12.5% NSF

1. Identification

Product Name: Sodium Hypochlorite, 12.5% NSF

Item #: MA0605008

Web SDS: S415

Synonyms: Chlorine Bleach

Recommended Use: Laboratory Reagent

Restrictions on Use: Any use other than recommended

Manufacturer:

ABC Chemical 999
Chemical Ln,
Nowhere, PA 98273
1-800-555-5555

In Case of Emergency:

Chemtrec US 1-800-424-9300
Chemtrec International 703-527-3887

2. Hazards Identification

OSHA Hazard Classification(s):

Corrosive to Metals - Category 1
Acute Toxicity - Oral - Category 4
Skin Corrosion - Category 1A Eye
Damage - Category 1

Signal Word: DANGER

Hazard Statement(s): May be corrosive to metals. Harmful if swallowed. Causes severe skin burns and eye damage. Causes serious eye damage. Very toxic to aquatic life.

Pictogram(s):



Precautionary Statement(s): Keep only in original container. Do not breathe mist, vapors, or spray. Wash hands, forearms, and exposed areas thoroughly after handling. Avoid release to the environment. Wear eye protection, face protection, protective clothing, protective gloves. IF SWALLOWED: Rinse mouth. Do NOT induce vomiting. IF ON SKIN(or hair): Take off all contaminated clothing immediately. Rinse skin with quick-drench shower. IF INHALED: Remove person to fresh air and keep at rest in a comfortable position for breathing. IF IN EYES: Rinse cautiously with water for several minutes. Remove contact lenses if present and easy to do. Continue rinsing. Immediately call a poison center, doctor, or physician. Wash contaminated clothing before reuse. Absorb spillage to prevent material damage. Store locked up. Dispose of contents and container in accordance to local, regional, national, or international regulations.

Descriptions of Hazards not otherwise classified: N/A

Percent of mixture with unknown acute toxicity: N/A

3. Composition and Information on Ingredients

Chemical Name	Common Name	CAS #	Concentration %
Sodium Hypochlorite	Chlorine Bleach	7681-52-9	12.5%

4. First Aid Measures

Eye Contact: Immediately flush with water for at least 15 minutes. Call a physician.

Skin Contact: Immediately flush with large quantities of cool water continuously for at least 15 minutes. Call a physician. Remove contaminated clothing and shoes. Do not put contaminated clothing and shoes back on. Wash clothing and shoes thoroughly in soap and water; rinse repeatedly in clean water and dry before reuse.

Inhalation: NOTE TO PHYSICIAN: Probable mucosal damage may contraindicate the use of gastric lavage. Move person to fresh air. If person is not breathing, call 911 or an ambulance. Call a poison control center or doctor for further treatment advice.

Ingestion: Call a poison control center or doctor immediately. Have the person sip a glass of water if able to swallow. Do not induce vomiting unless told to do so by a doctor or poison control center. Do not give anything by mouth to an unconscious person.

Symptoms: Contacted areas will exhibit severe irritation or burns. Burns may not be immediately apparent. Eye contact may cause permanent injury including blindness. If ingested, may cause nausea, vomiting and death. Effects may include circulatory collapse, delirium, coma and possible perforation of esophagus and stomach. May act as a sensitizer.

Recommendations for immediate medical care/special treatment: Get medical advice/attention if you feel unwell.

Appendix: Safety Data Sheet (SDS) Example

Sodium Hypochlorite, 12.5% NSF

5. Fire- Fighting Measures

Extinguishing Media: Dry chemical, carbon dioxide, alcohol foam.

Fire Hazards (Chemical): SPECIFIC HAZARDS ARISING FROM THE SUBSTANCE OR MIXTURE: Not considered flammable or explosive. Use extinguishing agents suitable for the surrounding fire and not contraindicated for use with sodium hypochlorite.

Sodium hypochlorite releases oxygen when heated, which may increase the severity of an existing fire. Use water spray to cool fire exposed surfaces and to protect personnel. Do not use a direct water stream

Special Protective Equipment: Wear self-contained breathing apparatus and full protective clothing. Use water spray to keep containers cool. Avoid inhalation of material or combustion by products. Firefighters should wear full protective clothing and NIOSH approved positive pressure self-contained breathing apparatus.

Precautions for Firefighters: Chlorine gas. Chlorine gas is an oxidizer and will support combustion.

6. Accidental Release Measures

Emergency Procedures: Evacuate the area of all unnecessary personnel. Wear suitable protective equipment. Eliminate all sources of ignition, provide ventilation, and use non-sparking tools.

Protective Equipment: See section 8

Environmental Precautions: Contain liquid spills with sand and absorb on inert material such as Hazorb or clay. Avoid breathing vapors. Ventilate area.

Containment and Clean-Up Procedures: Dispose with solid wastes. Avoid contact with acids. Do not discharge to sewers or waterways without proper treatment. Contact state and federal environment organizations for disposal.

7. Handling and Storage

Handling: Wear proper safety equipment when handling this product. Handle in accordance with good industrial hygiene and safety procedures. DO NOT MIX SODIUM HYPOCHLORITE 12.5% WITH ACIDS! THIS WILL FORM TOXIC CHLORINE GAS.

Storage: Store in a cool dry area, away from direct sunlight and heat to avoid deterioration. Keep containers closed when not in use. Vent container frequently and more often in hot weather to relieve pressure. Do not reuse or refill this container.

8. Exposure Controls/Personal Protection

OSHA Permissible Exposure Limits (PELs):

Reagent	CAS #	OSHA PEL TWA
Sodium hypochlorite	7681-52-9	N/a

ACGIH Threshold Limit Values (TLVs):

Reagent	CAS #	ACGIH PEL TLV	ACGIH STEL
Sodium Hypochlorite	7681-52-9	N/a	N/a

Engineering Controls: Use in a well ventilated area to prevent exposure. Maintain eyewash fountain and quick-drench facilities in work areas.

Personal Protective Measures: Wear gloves, lab coat, eye protection and impervious footwear. Contact lenses should not be worn when working with this material.

Special PPE Requirements: If ventilation hood not available wear respirator.

9. Physical and Chemical Properties Section

Appearance: N/A, Yellowish-green liquid

Molecular Weight: N/A

Molecular Formula: N/A

pH: 12.3-12.7

Boiling Point and Boiling Range: 284 F (140 C)

Melting Point/Freezing Point: -17 F (-27 C)

Appendix: Safety Data Sheet (SDS) Example

Sodium Hypochlorite, 12.5% NSF

Flash Point: N/A
Specific Gravity/Relative Density: 1.240
Odor: Mild Chlorine
Odor Threshold: 0.3 ppm (for chlorine)
Color: N/A
Flammability (solid/gas): N/A
Vapor Density: N/A
Upper/Lower flammability or explosive limits: N/A
Vapor Pressure: N/A
Evaporation Rate: N/A
Partition Coefficient: n-octanol/water: N/A
Viscosity: N/A
Auto-ignition temperature: N/A
Solubility: N/A
Decomposition Temperature: N/A

10. Stability and Reactivity

Reactivity: Will react with acids and ammonia to release toxic chlorine gas.
Chemical Stability: Stable
Conditions of Stability/Instability: Stable under recommended handling and storage conditions.
Stabilizers needed: None
Safety issue indicated by appearance change: N/A
Other: N/A
Hazardous Reactions: N/A
Hazardous Polymerization: Does not occur
Conditions to avoid: Direct sunlight and high temperatures (over 100 F)
Classes of Incompatible Materials: Acids, amines, ammonia compounds, oxidizing materials, peroxides, organic materials, reducing agents, cyanides, ethers, hydrocarbons, oxidizable materials, and most metals. DO NOT MIX WITH ACIDS, WILL FORM TOXIC CHLORINE GAS.
Hazardous Decomposition Products: Chlorine gas, hypochlorous acid, hydrochloric acid, oxygen gas, hydrogen chloride gas

11. Toxicological Information

Likely Routes of Exposure

Eyes: Causes severe eye damage.
Skin: Causes severe irritation.
Inhalation: May cause nasal and respiratory irritations
Ingestion: Harmful if swallowed

Signs or Symptoms of Exposure: Nausea.

Effects from short term exposure (delayed, immediate, chronic): Irritation to the eyes, nose, throat; headache, dizziness, nausea.

Acute Toxicity (Numerical Measures): LD50 Oral Rat: 8910 mg/kg LD50 Dermal Rabbit: >10,000 mg/m³ LC50 Inhalation: No data

Carcinogenicity (NTP, IARC, OSHA): This material contains no ingredient above de minimus concentrations known or suspected to cause cancer.

12. Ecological Information

Ecotoxicity: Freshwater Fish Data: LC50 (Bluegill sunfish): 2.90 mg/L/96 hours LC50 (Pimephales promelas): 1.40 mg/L/96 hours LC50 (Oncorhynchus mykiss): 0.90 mg/L/0.5 hours Invertebrate Toxicity Data: No data

Persistence and degradability: This product is inorganic and not subject to biodegradation.

Bioaccumulation Potential (octanol-water partition coefficient, BCF): N/A

Mobility in the soil: N/A

Adverse Environmental Effects: This product contains no hazardous air pollutants (HAPS).

APPENDIX: SDS FOR SODIUM HYPOCHLORITE

13. Disposal Considerations

Recommended Disposal Containers: Check with your local waste authorities*

Recommended Disposal Methods: Normal for sodium hypochlorite containing wastes. Sodium metabisulfite may be used to neutralize chlorine. May require pH adjustment for neutralization. Dispose in accordance with local, state and federal regulations. Do not discharge effluent containing this product to sewer systems without previously notifying the local sewage treatment plant authority. **Physical/Chemical Properties affecting Disposal:** See section 2 and section 9 applicable information.*

Special Precautions for Landfill and Incineration Activities: Check with your local waste authorities.*

Waste Stream: Consult your local or regional authorities.*

14. Transport Information

UN Number: UN1791

UN Proper Shipping Name: Hypochlorite solutions, 8, PG III, ERG# 154

Transport Hazard Class(es): 8

Packing Group Number: III

Environmental Hazards (IMDG code):

Marine Pollutant:

Transport in Bulk (IBC Code): N/A

Special Transport Precautions: N/A

15. Regulatory Information

OSHA: N/A

DOT: N/A

EPA: N/A

CPSC:

N/A

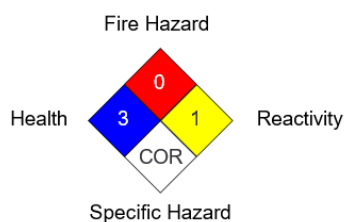
16. Other Information

Revision Date: 11/15/2016

NFPA

Health	3
Fire Hazard	0
Reactivity	1
Specific Hazard	COR

National Fire Protection Association (USA) NFPA



HMIS

Health	3
Flammability	0
Physical Hazard	1
Personal Protection	

Hazardous Material Information System HMIS

Health	3
Flammability	0
Physical Hazard	1
Personal Protection	

Notice to Reader:

To the best of our knowledge, the information contained herein is accurate. However, neither the above named supplier nor any of its subsidiaries assumes any liability whatsoever for the accuracy or completeness of the information contained herein.

APPENDIX: HYPOCHLORITE PRACTICE MATH PROBLEMS

Hypochlorite Practice Math Problems

- 1. In 24 hours, 4.2 gallons of 12% hypochlorite solution is fed. How much (in gallons) would you have to use if the concentration was 7%?**
 - a. 2.4 gallons
 - b. 5 gallons
 - c. 7.2 gallons
 - d. 10.1 gallons
- 2. In 10 hours, you feed 3.5 gallons of 12.5% hypochlorite solution. How many gallons would you have to use if the concentration is 6%?**
 - a. 7.3 gallons
 - b. 3.7 gallons
 - c. 1.4 gallons
 - d. 0.9 gallons
- 3. A plant flow is set at 2.2 MGD. The chlorine dose needs to be 2.0 mg/L. How many pounds of 12.5% sodium hypochlorite can the system expect to use each day?**
 - a. 294 pounds
 - b. 37 pounds
 - c. 0.3 pounds
 - d. 30 pounds
- 4. A system needs to determine how many pounds of 12.5% sodium hypochlorite they will use when the plant is set at a flow of 375,000 gpd. They need to maintain a chlorine dosage of 1.5 mg/L.**
 - a. 4.7 pounds
 - b. 47 pounds
 - c. 37.5 pounds
 - d. 3.8 pounds
- 5. A tank contains 575,000 gallons of water. This water is to receive a chlorine dose of 2.2 mg/L. How many pounds of calcium hypochlorite (65% available) will be required for this disinfection?**
 - a. 16.2 pounds
 - b. 10,550 pounds
 - c. 10.55 pounds
 - d. 0.162 pounds

APPENDIX: HYPOCHLORITE PRACTICE MATH PROBLEMS

6. Calculate the amount of chlorine required to dose an 800,000 gallon storage tank to a dose of 5 mg/L. You believe it is best to use granular calcium hypochlorite and the product information indicates it is 68% chlorine.
- a. 25 pounds
 - b. 33 pounds
 - c. 49 pounds
 - d. 60 pounds
7. After cleaning, a system needs to disinfect a 750,000 gallon storage tank. The system has decided on a dose of 25 mg/L. How many pounds of 68% calcium hypochlorite would they need to purchase for the job?
- a. 30 pounds
 - b. 75 pounds
 - c. 156 pounds
 - d. 230 pounds
8. How many pounds of chlorine would be needed to disinfect a 700,000 gallon tank that is $\frac{2}{3}$ full? It has been determined it needs dosed to 5 mg/l.
- a. 19.5 lbs
 - b. 29.2 lbs
 - c. 9.7 lbs
 - d. 39.5 lbs
9. A treatment plant uses sodium hypochlorite (12%) to disinfect the water. The target dose is 0.8 mg/L. They treat 250,000 gpd. How many pounds of sodium hypochlorite will they need to feed?
- a. 14 pounds
 - b. 10 pounds
 - c. 4.3 pounds
 - d. 1.7 pounds
10. A treatment plant uses 12.5% hypochlorite to disinfect the water. The required hypochlorite dosage is 2 mg/L and the plant flow is 300,000 gpd. How many gallons of 12.5% hypochlorite are required (12.5% hypo has 1.25 lbs/gal available chlorine)?
- a. 5.1 gallons
 - b. 4 gallons
 - c. 12 gallons
 - d. 400 gallons

APPENDIX: HYPOCHLORITE PRACTICE MATH PROBLEMS

- 11. The chlorine demand of the water is 1.4 mg/L. If the desired chlorine residual is 0.5 mg/L, what is the desired chlorine dose, in mg/L?**
- a. 0.9 mg/L
 - b. 1.3 mg/L
 - c. 1.5 mg/L
 - d. 1.9 mg/L
- 12. At a flow rate of 375 gpm, how many pounds of 67% calcium hypochlorite would be required to maintain a 0.8 mg/L chlorine residual in the finished water if the chlorine demand is 0.8 mg/L?**
- a. 3.6 pounds
 - b. 5.4 pounds
 - c. 7.2 pounds
 - d. 10.8 pounds
- 13. A system needs to maintain a chlorine residual of 0.8 mg/L. The chlorine demand is 1.2 mg/L and the plant flow is set at 500 gpm. How many pounds of 65% calcium hypochlorite would the system expect use in 8 hours?**
- a. 12 pounds
 - b. 6 pounds
 - c. 9 pounds
 - d. 11 pounds
- 14. How many gallons of 15% sodium hypochlorite (1.4 lbs/gal available chlorine) are required to treat 750,000 gpd with a desired chlorine residual of 0.8 mg/L and a demand of 0.6 mg/L?**
- a. 2 gallons
 - b. 4 gallons
 - c. 6 gallons
 - d. 8 gallons
- 15. How many gallons of 12½% sodium hypochlorite are required to treat 750,000 gpd with a desired residual of 1.2 mg/L and a chlorine demand of 0.5 mg/L? (note, 12½% has 1.2 lbs/gal available chlorine)**
- a. 7 gal/day
 - b. 9 gal/day
 - c. 11 gal/day
 - d. 13 gal/day

APPENDIX: HYPOCHLORITE PRACTICE MATH PROBLEMS

- 16. A plant is set at a flow rate of 3 MGD. Water enters into a clearwell that has a volume of 55,000 gallons. The chlorine residual of the outlet end of the tank is 1.4 mg/L. What is the CT in mg-min/L?**
- a. 25
 - b. 180
 - c. 37
 - d. 203
- 17. A free chlorine residual of 1.7 mg/L is measured at the end of the clearwell after 4 hours of detention time, what is the CT value in mg-min/L?**
- a. 6.8 mg-min/L
 - b. 80 mg-min/L
 - c. 240 mg-min/L
 - d. 408 mg-min/L
- 18. A plant is set at a flow rate of 2 MGD. Water enters into a clearwell that has a volume of 50,000 gallons. The chlorine residual of the outlet end of the tank is 0.9 mg/L. What is the CT in mg-min/L?**
- a. 11 mg-min/L
 - b. 22.5 mg-min/L
 - c. 32 mg-min/L
 - d. 44.5 mg-min/L
- 19. If the free chlorine residual of 1.8 mg/L is measured at the end of the clearwell after 3 hour of detention time, what is the CT value in mg-min/L?**
- a. 5 mg-min/L
 - b. 75 mg-min/L
 - c. 176 mg-min/L
 - d. 324 mg-min/L
- 20. If the free residual of 1.8 mg/L is measured at the entry point of the system, after 5 hours of detention time, what is the CT value in mg-min/L?**
- a. 16 mg-min/L
 - b. 95 mg-min/L
 - c. 275 mg-min/L
 - d. 540 mg-min/L

APPENDIX: HYPOCHLORITE PRACTICE MATH PROBLEMS

21. In 18 hours, you use 6 gallons of 15% sodium hypochlorite. How much (gallons) would you have to use if the concentration is 7%.

- a. 3 gallons
- b. 10 gallons
- c. 13 gallons
- d. 16 gallons

22. A system is using 12.5% sodium hypochlorite to disinfect at a dose of 1.5 mg/L. When the plant flow is set at 550,000 gpd, how many pounds of sodium hypochlorite should they expect to use?

- a. 7 pounds
- b. 35 pounds
- c. 55 pounds
- d. 155 pounds

A Module 25 answer key is available at the following website link:

<http://www.depweb.state.pa.us/operatorcenter/>

In the right-hand menu click "Training", then "DEP Training Modules"